Unit 1 Chemistry Study Guide

History of the Atom:

Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called atomos.

Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called hyle.

Antoine Lavoisier (1774):

Law of Conservation of Mass: Mass cannot be created or destroyed.

Joseph Proust (1779):

 Law of Definite Proportions: All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

John Dalton (1803 - 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first atomic theory.
- Daltons Atomic Theory:
 - All matter is composed of extremely small particles called atoms.
 - All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
 - Atoms cannot be subdivided, created, or destroyed.
 - Atoms to different elements can combine in simple whole number ratios to form chemical compounds.
 - o In chemical reactions, atoms are combined, separated or rearranged.
- Law of Multiple Proportions: Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model: ->



allow

- J.J. Thompson (1906):
 - Cathode Ray Experiment -
 - Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

Unit 6 Chemistry Study Guide

Susan McMurry

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