

Unit 3 Worksheet 4 – Quantitative Energy Problems Part 2

Energy constants (H₂O)

334 J/g	Heat of fusion (melting or freezing) H_f
2260 J/g	Heat of vaporization (evaporating or condensing) H_v
2.1 J/g°C	Heat capacity (c) of solid water
4.18 J/g°C	Heat capacity (c) of liquid water

For each of the problems sketch a warming or cooling curve to help you decide which equation(s) to use to solve the problem. Keep a reasonable number of sig figs in your answers.

1. How much energy must be absorbed by a 150 g sample of ice at 0.0 °C that melts and then warms to 25.0°C?
2. Suppose in the Icy Hot lab that the burner transfers 325 kJ of energy to 450 g of liquid water at 20.°C. What mass of the water would be boiled away?
3. A 12oz can of soft drink (assume $m = 340$ g) at 25°C is placed in a freezer where the temperature is - 12 °C. How much energy must be removed from the soft drink for it to reach this temperature?

Unit 3 4 Quantitative Energy Problems Answers

John Michels



Unit 3 4 Quantitative Energy Problems Answers:

Jacaranda Live It Up 2 VCE Physical Education Units 3&4, 5e learnON & Print Sally Nelson,Kirsty Walsh,Paul Stockdale,Andrew Hamilton,James W. Guthrie,Sam Millar,2024-11-12 Created in association with ACPHER a team of authors provide their teaching and examination expertise to ensure the delivery of a comprehensive and highly supportive resource Practical exercises are integrated throughout this resource to facilitate the application of theoretical knowledge through practical activity Key skill subtopics with videos break down what is expected for each key skill and provide an annotated model answer to help student understanding Scaffolded extended response subtopics with videos that unpack how to answer this difficult style of exam question step by step Exam tip boxes highlight how to avoid common pitfalls and tackle questions Immediate feedback for every question online with targeted question sets for every key knowledge point including quick quiz questions ACPHER trial exam questions and past VCAA exam questions **The Quantum in Chemistry** Roger

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