UNIT 2 - BIOCHEMISTRY TEST STUDY GUIDE

Basic Chemistry

Use the terms below to complete the statements below. (HINT: Some of the words will be used in multiple statements.)

Electrons Isotopes
Electron cloud Neutral
Ion Neutrons

Proton 1

The two regions of the Atom are the nucleus and the electron cloud.

The particles found within an atom are the proton, neutron and the electron.

The mass number of an element is equal to the protons and neutrons. The mass of an atom is located in the nucleus.

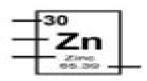
Since ATOMS are neutral the protons and the electrons will always be equal. If they weren't, then there would be a charge and the atom would be called an ion.

Atoms that have the same number of protons, but have different numbers of neutrons are called isotope.

Complete the table below describing the particles in the atom

Particle	Abbreviation	Location	Charge	
Proton	P+	Nucleus		
Neutron	No	Neucleus	No charge	
Electron	E.	Electron cloud	Negative	

Label the periodic table block below with the following terms: element name, element symbol, average atomic mass, & atomic number.



What information does the atomic number tell you about the atom?

The number of protons in the atom.

What information does a mass number tell you about the atom?

The number of protons and neutrons in the atom

How can you calculate the number of neutrons in an atom? Atomic mass – atomic number

List the names and symbols of the elements described below:

Elements that make up 96% of living matter

Element Name	Received .	Hydrogen	Element Symbol	Oxygen	Element Namedian	Nitrogen	Darrier .
Carpon			- 14		0		Pul

What are the other 4% of elements called? Minerals and Trace elements

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Katherine J. Denniston, Denniston Katherine

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Table of Contents Unit 2 Biochemistry Study Guide

- 1. Understanding the eBook Unit 2 Biochemistry Study Guide
 - The Rise of Digital Reading Unit 2 Biochemistry Study Guide
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Unit 2 Biochemistry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 2 Biochemistry Study Guide
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Unit 2 Biochemistry Study Guide
 - Personalized Recommendations
 - Unit 2 Biochemistry Study Guide User Reviews and Ratings
 - Unit 2 Biochemistry Study Guide and Bestseller Lists
- 5. Accessing Unit 2 Biochemistry Study Guide Free and Paid eBooks

- Unit 2 Biochemistry Study Guide Public Domain eBooks
- Unit 2 Biochemistry Study Guide eBook Subscription Services
- Unit 2 Biochemistry Study Guide Budget-Friendly Options
- 6. Navigating Unit 2 Biochemistry Study Guide eBook Formats
 - o ePub, PDF, MOBI, and More
 - Unit 2 Biochemistry Study Guide Compatibility with Devices
 - Unit 2 Biochemistry Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Unit 2 Biochemistry Study Guide
 - Highlighting and Note-Taking Unit 2 Biochemistry Study Guide
 - Interactive Elements Unit 2 Biochemistry Study Guide
- 8. Staying Engaged with Unit 2 Biochemistry Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 2 Biochemistry Study Guide
- 9. Balancing eBooks and Physical Books Unit 2 Biochemistry Study Guide
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Unit 2 Biochemistry Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 2 Biochemistry Study Guide
 - Setting Reading Goals Unit 2 Biochemistry Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 2 Biochemistry Study Guide
 - Fact-Checking eBook Content of Unit 2 Biochemistry Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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