

Chapter 1 - Structure and Bonding (test bank) KEY

1. What is the ground-state electronic configuration of a carbon atom?

- A. $1s^2, 2s^2, 2p^5$
- B. $1s^2, 2s^2, 2p^2$**
- C. $1s^2, 2s^2, 2p^6$
- D. $1s^2, 2s^2, 2p^4$

Accessibility: Keyboard Navigation

Bloom's Level: 2. Understand

Difficulty: Easy

Gradable: automatic

Section: 01.01

Subtopic: Periodic table trends

Topic: Structure and Bonding

2. What is the ground-state electronic configuration of a fluorine atom?

- A. $1s^2, 2s^2, 2p^2$
- B. $1s^2, 2s^2, 2p^3$**
- C. $1s^2, 2s^2, 2p^4$
- D. $1s^2, 2s^2, 2p^5$

Accessibility: Keyboard Navigation

Bloom's Level: 2. Understand

Difficulty: Easy

Gradable: automatic

Section: 01.01

Subtopic: Periodic table trends

Topic: Structure and Bonding

3. What is the ground-state electronic configuration of a magnesium cation (Mg^{2+})?

- A. $1s^2, 2s^2, 2p^6$**
- B. $1s^2, 2s^2, 2p^6, 3s^1$
- C. $1s^2, 2s^2, 2p^6, 3s^2$
- D. $1s^2, 2s^2, 2p^6, 3s^2, 3p^2$

Accessibility: Keyboard Navigation

Bloom's Level: 2. Understand

Difficulty: Easy

Gradable: automatic

Section: 01.01

Subtopic: Periodic table trends

Topic: Structure and Bonding

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