

Straighterline Chemistry Final- deck 3

Which is more soluble in an acidic solution than in pure water?

Select one:

- a. CuI
- b. PbCl_2
- c. $\text{Ca}_3(\text{PO}_4)_2$
- d. NaNO_3
- e. NaBr - ANSc. $\text{Ca}_3(\text{PO}_4)_2$

What is the name of the principle of selective precipitation used to identify the types of ions present in a solution?

Select one:

- a. Ionization
- b. Selective ion precipitation
- c. Selective ion typing
- d. Limited precipitation
- e. Qualitative analysis - ANSe. Qualitative analysis

A 10.0-mL sample of 0.75 M $\text{CH}_3\text{CH}_2\text{COOH}$ is titrated with 0.30 M NaOH . What is the pH of the solution after 22.0 mL of NaOH have been added to the acid?

$[\text{Ka}(\text{CH}_3\text{CH}_2\text{COOH}) = 1.3 \times 10^{-5}]$

Select one:

- a. 5.75
- b. 4.94
- c. 4.83
- d. 4.02
- e. 3.95 - ANSa. 5.75

Methyl red is a common acid-base indicator. It has a Ka equal to 6.3×10^{-6} . Its un-ionized form is red and its anionic form is yellow. What color would a methyl red solution have at $\text{pH} = 7.8$?

Select one:

- a. green
- b. red
- c. blue
- d. yellow
- e. violet - ANSd. yellow

A 25.0-mL sample of 0.10 M $\text{C}_2\text{H}_5\text{NH}_2$ (ethylamine) is titrated with 0.15 M HCl . What is the pH of the solution after 9.00 mL of acid have been added to the amine?

$[\text{Kb}(\text{C}_2\text{H}_5\text{NH}_2) = 6.5 \times 10^{-4}]$

Select one:

Straighterline Chemistry Answer Key

Daniel C. Tofan



Straighterline Chemistry Answer Key:

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General Chemistry Workbook Daniel C. Tofan,2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

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