

Naming & Writing Ionic Formulas III

Name: _____

Period: _____ Date: _____

I. In the space before the formula write I for ionic compounds, IH for hydrates, and A for acids. Then, write the name of the compound.

- | | |
|---|--|
| ____ 1. HCl _____ | ____ 16. Ba(BrO ₃) ₂ _____ |
| ____ 2. KCl _____ | ____ 17. NaC ₂ H ₃ O ₂ _____ |
| ____ 3. NH ₄ OH _____ | ____ 18. HClO ₄ _____ |
| ____ 4. MgSO ₄ •7H ₂ O _____ | ____ 19. Na ₂ SO ₃ _____ |
| ____ 5. HNO ₃ _____ | ____ 20. LiH _____ |
| ____ 6. Cu ₂ O _____ | ____ 21. MgBr ₂ _____ |
| ____ 7. NaOH _____ | ____ 22. H ₂ CO ₃ _____ |
| ____ 8. HF _____ | ____ 23. Ba ₃ As ₂ _____ |
| ____ 9. NaHCO ₃ _____ | ____ 24. Pb(OH) ₂ _____ |
| ____ 10. CuSO ₄ •5H ₂ O _____ | ____ 25. NaClO _____ |
| ____ 11. Zn(NO ₃) ₂ _____ | ____ 26. HC ₂ H ₃ O ₂ _____ |
| ____ 12. CsOH _____ | ____ 27. Cu ₃ P _____ |
| ____ 13. SnCl ₂ •2H ₂ O _____ | ____ 28. K ₂ C ₂ O ₄ _____ |
| ____ 14. H ₂ SO ₄ _____ | ____ 29. H ₃ C ₆ H ₅ O ₇ _____ |
| ____ 15. FeCO ₃ _____ | ____ 30. Ca(MnO ₄) ₂ _____ |

II. Write the formula for the following compounds:

- | | |
|--|--|
| 31. manganese (II) hydroxide _____ | 46. cobalt (II) chloride hexahydrate _____ |
| 32. silver chloride _____ | 47. aluminum oxide _____ |
| 33. ammonium sulfide _____ | 48. lead (IV) bicarbonate _____ |
| 34. phosphoric acid _____ | 49. strontium oxalate _____ |
| 35. barium hydroxide octahydrate _____ | 50. rubidium perchlorate _____ |
| 36. cesium carbonate _____ | 51. chlorous acid _____ |
| 37. chromium (III) phosphide _____ | 52. ammonium sulfate monohydrate _____ |
| 38. cobalt (III) nitrite _____ | 53. titanium (IV) carbonate _____ |
| 39. sulfurous acid _____ | 54. nitric acid _____ |
| 40. hydrobromic acid _____ | 55. ammonium bromate _____ |
| 41. lead (IV) chloride _____ | 56. perchloric acid _____ |
| 42. cadmium phosphate _____ | 57. strontium hydroxide _____ |
| 43. nickel (II) chloride _____ | 58. manganese (II) sulfide _____ |
| 44. vanadium (III) dichromate _____ | 59. vanadium (V) phosphate _____ |
| 45. gold (III) bromide _____ | 60. zinc iodide _____ |

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Michal Rosen-Zvi



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