

Name: KEY

Section: _____

**Chapter 7 Practice Worksheet:
Covalent Bonds and Molecular Structure**

1) How are ionic bonds and covalent bonds different?

Ionic bonds result from the transfer of electrons from one atom to another; Covalent bonds result from two atoms sharing electrons.

2) Describe the relationship between the length of a bond and the strength of that bond.

Strength of a bond increases as the bond gets shorter (inverse relationship)

3) Identify the type(s) of bond(s) found in the following molecules:

- a. CCl_4 covalent
- b. Li_2O ionic
- c. NF_3 covalent
- d. CaSO_4 ionic and covalent
- e. SO_2 covalent
- f. $\text{Mg}(\text{OH})_2$ ionic and covalent

4) Define electronegativity.

Electronegativity is the ability for an atom in a molecule to attract electrons to itself.

5) Use electronegativity values to place the following elements in **increasing** order: F, N, H, P, Si, C, O

$\text{Si} < \text{C} \approx \text{P} \approx \text{H} < \text{N} < \text{O} < \text{F}$

6) Determine if the bond between atoms in each example below is nonpolar covalent, polar covalent, or ionic.

- | | | | |
|------------------|--------------|-----------------|--------------|
| a. H_2 | <u>npc</u> | e. NF | <u>pc</u> |
| b. PCl | <u>pc</u> | f. MgO | <u>ionic</u> |
| c. F_2 | <u>npc</u> | g. CH | <u>npc</u> |
| d. NaBr | <u>ionic</u> | h. HCl | <u>pc</u> |

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Answers

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Table of Contents Using Science Skills Relating Concepts Chapter 7 Answers

1. Understanding the eBook Using Science Skills Relating Concepts Chapter 7 Answers
 - The Rise of Digital Reading Using Science Skills Relating Concepts Chapter 7 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Using Science Skills Relating Concepts Chapter 7 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Using Science Skills Relating Concepts Chapter 7 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Using Science Skills Relating Concepts Chapter 7 Answers
 - Personalized Recommendations
 - Using Science Skills Relating Concepts Chapter 7 Answers User Reviews and Ratings
 - Using Science Skills Relating Concepts Chapter 7 Answers and Bestseller Lists
5. Accessing Using Science Skills Relating Concepts Chapter 7 Answers Free and Paid eBooks

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- 6. Navigating Using Science Skills Relating Concepts Chapter 7 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Using Science Skills Relating Concepts Chapter 7 Answers Compatibility with Devices
 - Using Science Skills Relating Concepts Chapter 7 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Using Science Skills Relating Concepts Chapter 7 Answers
 - Highlighting and Note-Taking Using Science Skills Relating Concepts Chapter 7 Answers
 - Interactive Elements Using Science Skills Relating Concepts Chapter 7 Answers
- 8. Staying Engaged with Using Science Skills Relating Concepts Chapter 7 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Using Science Skills Relating Concepts Chapter 7 Answers
- 9. Balancing eBooks and Physical Books Using Science Skills Relating Concepts Chapter 7 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Using Science Skills Relating Concepts Chapter 7 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Using Science Skills Relating Concepts Chapter 7 Answers
 - Setting Reading Goals Using Science Skills Relating Concepts Chapter 7 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Using Science Skills Relating Concepts Chapter 7 Answers
 - Fact-Checking eBook Content of Using Science Skills Relating Concepts Chapter 7 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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