

SYBEX Sample Chapter

CCNA™ : Cisco® Certified Network Associate Study Guide, 5th Edition (640-801)

Todd Lammle

Chapter 3: IP Subnetting and Variable Length Subnet Masks (VLSMs)

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ISBN: 0-7821-4391-1

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l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} \text{ s } \text{MnO}_4^{-}$
4 a a q is

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow \text{AuCN} + 2\text{e}^- + 2\text{H}^+ + \text{O}_2$ $\text{H}_2\text{O}_2 + 2\text{H}^+ + 2\text{e}^- \rightarrow 2\text{H}_2\text{O}$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} + 2\text{AuCN} + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow 2\text{OH}^- + \text{H}_2\text{O}_2$ the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h and

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