



Ordinary Thinking

Objective Questions

Oxidation, Reduction

- H_2O_2 reduces MnO_4^- ion to [KCET (Med.) 2000]
 - Mn^{2+}
 - Mn^{3+}
 - Mn^{2+}
 - Mn^{+}
- When a sulphur atom becomes a sulphide ion [AMU 1999]
 - There is no change in the composition of atom
 - It gains two electrons
 - The mass number changes
 - None of these
- The ultimate products of oxidation of most of hydrogen and carbon in food stuffs are [DCE 2001]
 - H_2O alone
 - CO_2 alone
 - H_2O and CO_2
 - None of these
- When P reacts with caustic soda, the products are PH_3 and NaH_2PO_2 . This reaction is an example of [IIT 1980; Kurukshetra CEE 1993; CPMT 1997]
 - Oxidation
 - Reduction
 - Oxidation and reduction (Redox)
 - Neutralization
- Which one of the following does not get oxidised by bromine water [MP PET/PMT 1985]
 - Fe^{2+} to Fe^{3+}
 - Cl^{-} to Cl^{+}
 - Mn^{2+} to MnO_4^-
 - S_n^{2-} to S_n^{4-}
- In the reaction $H_2S + NO_3^- \rightarrow H_2O + NO + S$, H_2S is
 - Oxidised
 - Reduced
 - Precipitated
 - None of these
- The conversion of FeO_2 to $Fe(NO_3)_2$ is
 - Oxidation
 - Reduction
 - Neither oxidation nor reduction
 - Both oxidation and reduction
- In the course of a chemical reaction an oxidant [MP PMT 1986]
 - Loses electrons
 - Gains electrons
 - Both loses and gains electron
 - Electron change takes place
- $2Cu \rightarrow Cu + Cu_2$, the reaction is [RPMT 1997]
 - Redox
 - Neutralisation
 - Oxidation
 - Reduction
- H_2S reacts with halogens, the halogens [IPMER 2000]
 - Form sulphur halides
 - Are oxidised
 - Are reduced
 - None of these
- H_2O_2 reduces $K_2Fe(CN)_6$ [MP PMT 1985]
 - In neutral solution
 - In acidic solution
 - In non-polar solvent
 - In alkaline solution
- Max. number of moles of electrons taken up by one mole of NO_2^- when it is reduced to [DPMT 2002]
 - NH_3
 - NH_4OH
 - NO
 - NO_2
- In the reaction $3Mg + N_2 \rightarrow Mg_3N_2$ [MP PMT 1999]
 - Magnesium is reduced
 - Magnesium is oxidized
 - Nitrogen is oxidized
 - None of these
- When sodium metal is dissolved in liquid ammonia, blue colour solution is formed. The blue colour is due to [NCERT 1981]
 - Solvated Na^+ ions
 - Solvated electrons
 - Solvated NH_4^+ ions
 - Solvated protons
- Following reaction describes the rusting of iron $4Fe + 3O_2 \rightarrow 4Fe^{2+} + 6O^{2-}$

Which one of the following statement is incorrect [NCERT 1981; MNR 1991; AIIMS 1995]

 - This is an example of a redox reaction
 - Metallic iron is reduced to Fe^{2+}
 - Fe^{2+} is an oxidising agent
 - Metallic iron is a reducing agent
- $SrCl_2$ gives a precipitate with a solution of $HgCl_2$. In this process $HgCl_2$ is [CPMT 1983]
 - Reduced
 - Oxidised
 - Converted into a complex compound containing both Sr and Hg
 - Converted into a chloro complex of Hg
- Oxidation involves [NCERT 1971, 81; CPMT 1980, 82, 83; MP PMT 1983]
 - Loss of electrons
 - Gain of electrons
 - Increase in the valency of negative part
 - Decrease in the valency of positive part
- Incorrect statement regarding rusting is [MP PET 2000]
 - Metallic iron is oxidised to Fe^{2+} ions
 - Metallic iron is reduced to Fe^{2+} ions
 - Oxygen gas is reduced to oxide ion
 - Yellowish – brown product is formed
- When copper turnings are added to silver nitrate solution, a blue coloured solution is formed after some time. It is because, copper [CPMT 1974, 79; DPMT 200]
 - Displaces silver from the solution

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Wolfgang Guggemos



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