

Gas Laws Study Guide

Write Gas Laws Principles, identify the units, and complete the blanks in the following table.

1. A sample of gas is at 1.0 atm, 27°C, and 1.0 L.

2. The pressure is doubled and the volume is halved.

3. A sample of gas is at 1.0 atm, 27°C, and 1.0 L.

4. A sample of gas is at 1.0 atm, 27°C, and 1.0 L.

5. A sample of gas is at 1.0 atm, 27°C, and 1.0 L.

6. A sample of gas is at 1.0 atm, 27°C, and 1.0 L.

Name: _____ Date: _____

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Standard Temperature = _____ °C = _____ K

Standard Pressure = _____ atm = _____ mmHg = _____ kPa = _____ psi

Standard Volume = _____ L = _____ mL

Write the formula for the Gas Laws.

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Combined Gas Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Charles' Law

$$P_1 V_1 = P_2 V_2$$

Boyle's Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Gay-Lussac's Law

$$PV = nRT$$

Ideal Gas Law

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

Avogadro's Law

Complete the following statements by filling in the blanks & circling the correct words.

As Pressure increases (increases/decreases) and Volume is constant, meaning that they have a direct/inverse relationship.

As Volume increases Temperature (increases/decreases) given that _____ is constant meaning that they have a direct/inverse relationship.

As Pressure increases Temperature (increases/decreases) given that _____ is constant meaning that they have a direct/inverse relationship.

As the number of moles of gas in a container increases the Volume (increases/decreases), the Pressure (increases/decreases), and the Temperature (increases/decreases).

Study Guide For Content Mastery Gas Laws

LP Steffe



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