

Inorganic Chemistry

What is inorganic chemistry? All of the chemistry not covered in carbon-centered organic and biochemistry—the remaining 90 or so elements!

PERIODIC TABLE OF THE ELEMENTS

Periodic Table

1. Elemental Forms

- Monatomic** - Noble Gases: helium, neon, argon, krypton, xenon, radon; non-reactive
- Diatomic molecules**
Gas: fluorine, chlorine, hydrogen, oxygen, nitrogen
Liquid: bromine Solid: iodine
- Molecular solid**: sulfur
- Covalent solid**: carbon (diamond), graphite
- Metals**: most common type of element
- High coordination numbers; very dense
- Delocalized electrons form energy bands

2. Types of Compounds: large variety due to the "ionic-covalent-metallic" range

- Ionic**: electron exchange
- Covalent**: shared electrons
- Polar covalent**: unequal sharing
- Ion complex**: central ion + solvent
- Transition metal compounds**
- Coordination compound: non-bonding ligands
- Organometallic: metal-C bond
- Complex inorganic**: metal oxides and silicates

3. Periodicity: chemical properties of an element depend on valence electrons Family of elements: columns in Periodic Table

Periodic Trends

- Chemistry determined by **valence electrons**
- Atoms or ions get a bit larger as you go down the column; electronic shells are added
- Cation and anion **charge** decrease

REACTIONS

Inorganic = many types of reactions

- Acid-base neutralization**
- Dissolution of electrolytes in water**
- Oxidation**: loss of electrons, add oxygen
- Reduction**: gain of electrons, remove oxygen

TIP: Remember the acronym, OILRIG

- Precipitation**: dissolved ions form solids
- Hydrolysis**: water reacts with a compound
- Decomposition** of compounds to elements

CHEMISTRY OF THE METALS

1. Group 1: Alkali Metals

- Configuration**: noble gas + 1 electron, only valence: +1
- Natural occurrence**: ocean salts, oxide minerals
- Common magnetic**: strong base (NaOH, KOH); ionic salts (NaCl, KCl); soluble carbonate, nitrate, sulfate
- Reactions**: pure metals react violently with water; never found in nature as pure metal
 $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH(aq)} + \text{H}_2\text{(g)}$
- Spontaneous electrolysis** of molten salts
 $\text{NaCl(l)} \rightarrow \text{Na(l)} + \text{Cl}_2\text{(g)}$
- Solubility**: water forms cage $\text{M} \cdot \text{H}_2\text{O}$
Hydration: H^+ (2-3) to K^+ (6-8)
Hydration energy: varies from low and weak

Li
Na
K
Rb
Cs

2. Group 2: Alkaline Earth Metals

- Configuration**: filled s-subshell, only valence: +2
- Natural occurrence**: mineral carbonates, silicates, ocean salts
- Reactions**:
 $\text{Ca(s)} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2\text{(aq)} + \text{H}_2\text{(g)}$

Be
Mg
Ca
Sr
Ba

3. 1st Transition Series: Ti-Cu

Ti	V	Cr	Mn	Fe	Co	Ni	Cu
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- Configuration**: partially filled d-subshell; diverse chemical and electronic properties
- Periodic Trends**: electronegativity fairly constant; electrons obey Hund's Rule; density increases with Z
- Ligand Field Effect**: split d-orbitals stabilize

APP: white paint pigment



APP: attach organic groups to inorganic solids

EX: Manganese Chemistry

Valence ranges from 0 to 7, d⁵ to d⁰

Potassiumate: MnO_4^- , oxidizing agent

Mn(VII): easily reduced

4. 2nd & 3rd Transition Series: Zr-Au

Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag
Hf	Ta	W	Re	Os	Ir	Pt	Au

- Configuration**: comparable to the 1st transition series
- Trends in radii**: "lanthanide contraction"
- minimal change in radii for large Z; weak screening from f-orbitals
- M-M multiple bonds**: 2nd and 3rd series cluster compounds; delta (δ) bond, face-face d-orbitals

delta bond



EX: Tetraammine of Mn, Rh & Ru

d. **Complex molecular structure**: $\text{M(NH}_3)_4$ exists as a tetramer

#47 Silver: normally a non-reactive metal
Silver reacts with S $\rightarrow$ Ag_2S - silver tarnish
 $\text{Ag}_2\text{O} + \text{heat} \rightarrow \text{Ag}$

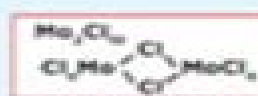
#42 Molybdenum

Quadruple M-M bond

Molybdenum $\text{Mo}_2\text{Cl}_8^{4-}$

tetrahedral - no lone pairs

Hallmark: MoF_6 , dimer - Mo_2F_{10}



#43 Technetium: synthetic element

5. Platinum Metals: Ru, Os, Rh, Ir, Pd, Pt

- Oxidation state** (0-6); oxygenation platinum
APP: catalysts for hydrogenation; catalytic converter

6. Scandium, Yttrium, Lanthanum & the Lanthanides

- Lanthanide element**: with occupied f-subshell

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Jackson,2010-05-31 Our 3 panel 6 page guide is jam packed with chemistry information not covered in organic chemistry and biochemistry Perfect for science students who want to focus on the aspects of chemistry that go beyond what is covered in material that deals with organic topics this guide covers the key concepts principles figures and formulas that inorganic chemistry students will need to know in order to succeed Tables images and graphic elements further enhance the text

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