



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol $\rightleftharpoons$.

What is Equilibrium? Equilibrium is a state of balance between two or more opposing processes. In chemistry, it refers to a reaction where the forward and reverse rates are equal, and the concentrations of reactants and products remain constant over time. Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the net change is zero.	Equilibrium Constant The equilibrium constant (K) is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant is given by: $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Graph showing concentration vs. time for a reaction at equilibrium. Reactant concentration: $[A]_0 - [A]_t = [A]_{eq}$ Product concentration: $[C]_0 + [C]_t = [C]_{eq}$
Equilibrium Constant, K_c K_c is the equilibrium constant expressed in terms of molar concentrations. It is used for reactions in the liquid phase. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant K_c is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Equilibrium Constant, K_p K_p is the equilibrium constant expressed in terms of partial pressures. It is used for reactions in the gas phase. For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant K_p is given by: $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	
Equilibrium Constant, K K is the general equilibrium constant, which can be expressed in terms of either concentrations (K_c) or partial pressures (K_p). For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant K is given by: $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	Relationship between K_c and K_p The relationship between K_c and K_p is given by the equation: $K_p = K_c (RT)^{\Delta n}$ where R is the gas constant, T is the temperature in Kelvin, and Δn is the change in the number of moles of gas.	
Le Chatelier's Principle: the effect of changing conditions on equilibrium		
Concentration	Increasing the concentration of a reactant shifts the equilibrium to the right (towards the products). Increasing the concentration of a product shifts the equilibrium to the left (towards the reactants).	For the reaction: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing the concentration of N_2 or H_2 will shift the equilibrium to the right, producing more NH_3.
Pressure	Increasing the pressure shifts the equilibrium towards the side with fewer moles of gas. Decreasing the pressure shifts the equilibrium towards the side with more moles of gas.	For the reaction: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing the pressure will shift the equilibrium to the right, as there are fewer moles of gas on the product side.
Temperature	Increasing the temperature shifts the equilibrium towards the endothermic direction. Decreasing the temperature shifts the equilibrium towards the exothermic direction.	For the reaction: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ This reaction is exothermic. Increasing the temperature will shift the equilibrium to the left, as the reverse reaction is endothermic.

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