

CHAPTER 14 STUDY GUIDE

Mixtures and Solutions

Section 14.1 Heterogeneous and Homogeneous Mixtures

In your textbook, read about suspensions and colloids.

For each statement below, write *true* or *false*.

- | | |
|--------------|---|
| <u>false</u> | 1. A solution is a mixture containing particles that settle out of the mixture if left undisturbed. |
| <u>true</u> | 2. The most abundant substance in a colloid is the dispersion medium. |
| <u>false</u> | 3. A colloid can be separated by filtration. |
| <u>true</u> | 4. A solid emulsion consists of a liquid dispersed in a solid. |
| <u>true</u> | 5. Whipped cream is an example of a foam. |
| <u>false</u> | 6. In an aerosol, the dispersing medium is a liquid. |
| <u>true</u> | 7. Brownian motion results from the collisions of particles of the dispersion medium with the dispersed particles. |
| <u>true</u> | 8. Dispersed particles in a colloid do not tend to settle out because they have polar or charged atomic groups on their surfaces. |
| <u>false</u> | 9. Stirring an electrolyte into a colloid stabilizes the colloid. |
| <u>true</u> | 10. Colloids demonstrate the Tyndall effect. |

The table below lists the characteristics of particles in colloids, solutions, and suspensions. Place a check in the column of each mixture whose particles have a particular characteristic.

Characteristics of Particles	Colloid	Solution	Suspension
11. Less than 1 nm in diameter		☒	☒
12. Between 1 nm and 1000 nm in diameter	☒	☒	☒
13. More than 1000 nm in diameter	☒	☒	
14. Settle out if undisturbed			
15. Pass through standard filter paper	☒		
16. Lower vapor pressure			☒
17. Scatter light	☒		

Study Guide For Chapter 2 substances Mixtures And Solubility

M Walker



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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

General, Organic, and Biochemistry Study Guide Marcia L. Gillette, Ira Blei, Wendy Gloffke, George Odian, 2006-01-27 This study guide provides reader friendly reinforcement of the concepts covered in the textbook Features include Chapter outlines Are you able to Worked text problems Fill ins Test yourself Concept maps Can also be used for Blei and Odian's Organic and Biochemistry

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Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter

Study Guide to Accompany Chemical Principles, Properties, and Reactions Kenneth L. Henold, 1984

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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 Student Study Guide Fundamentals of Chemistry by David E. Goldberg Margaret G. Kimble, 1994 *Study Guide to Accompany Basics for Chemistry* Martha Mackin, 1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple

Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students *Matter* Prentice-Hall Staff,1994

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Table of Contents Study Guide For Chapter 2substances Mixtures And Solubility

1. Understanding the eBook Study Guide For Chapter 2substances Mixtures And Solubility
 - The Rise of Digital Reading Study Guide For Chapter 2substances Mixtures And Solubility
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Chapter 2substances Mixtures And Solubility
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Chapter 2substances Mixtures And Solubility
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Chapter 2substances Mixtures And Solubility
 - Personalized Recommendations
 - Study Guide For Chapter 2substances Mixtures And Solubility User Reviews and Ratings
 - Study Guide For Chapter 2substances Mixtures And Solubility and Bestseller Lists
5. Accessing Study Guide For Chapter 2substances Mixtures And Solubility Free and Paid eBooks
 - Study Guide For Chapter 2substances Mixtures And Solubility Public Domain eBooks
 - Study Guide For Chapter 2substances Mixtures And Solubility eBook Subscription Services
 - Study Guide For Chapter 2substances Mixtures And Solubility Budget-Friendly Options
6. Navigating Study Guide For Chapter 2substances Mixtures And Solubility eBook Formats

- ePub, PDF, MOBI, and More
 - Study Guide For Chapter 2substances Mixtures And Solubility Compatibility with Devices
 - Study Guide For Chapter 2substances Mixtures And Solubility Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide For Chapter 2substances Mixtures And Solubility
 - Highlighting and Note-Taking Study Guide For Chapter 2substances Mixtures And Solubility
 - Interactive Elements Study Guide For Chapter 2substances Mixtures And Solubility
 8. Staying Engaged with Study Guide For Chapter 2substances Mixtures And Solubility
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Chapter 2substances Mixtures And Solubility
 9. Balancing eBooks and Physical Books Study Guide For Chapter 2substances Mixtures And Solubility
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide For Chapter 2substances Mixtures And Solubility
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Study Guide For Chapter 2substances Mixtures And Solubility
 - Setting Reading Goals Study Guide For Chapter 2substances Mixtures And Solubility
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Study Guide For Chapter 2substances Mixtures And Solubility
 - Fact-Checking eBook Content of Study Guide For Chapter 2substances Mixtures And Solubility
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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