

Unit Two Chemistry Exam

Multiple Choice 20 marks

You need to record your answers in the appropriate section in the ANSWER BOOKLET.

Question 1

When ethanol dissolves in water the bonds formed between the ethanol and water molecules are known as

- A. dispersion forces
- B. ion-dipole attraction
- C. hydrogen bonds
- D. ionisation

Question 2

When NaOH dissociates in water, the equation that describes the process is

- A. $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{(aq)} + \text{OH}^- \text{(aq)}$
- B. $\text{NaOH (s)} + \text{H}_2\text{O (l)} \rightarrow \text{NaO}^- \text{(aq)} + \text{H}_2\text{O}^+ \text{(aq)}$
- C. $\text{NaOH (s)} + \text{H}_2\text{O (l)} \rightarrow \text{NaOH}_2^+ \text{(aq)} + \text{OH}^- \text{(aq)}$
- D. $\text{NaOH (aq)} + \text{HCl (aq)} \rightarrow \text{NaCl (s)} + \text{H}_2\text{O (l)}$

Question 3

How much heat energy is required to raise the temperature of 250 mL of water from 17 °C to 100 °C? Assume that 1 mL of water has a mass of 1 g and that the specific heat capacity of water is $4.2 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$.

- A. 8.7 J
- B. 87,150 KJ
- C. 8.7 KJ
- D. $8.7 \times 10^4 \text{ J}$

Question 4

A solution containing 0.25 g of NaCl dissolved in 50 mL of deionised water will have a concentration of

- A. 5.0 gL^{-1}
- B. 4.3 M
- C. $5.0 \times 10^{-3} \text{ mgL}^{-1}$
- D. $2.5 \times 10^{-4} \text{ mol/L}$

Question 5

A solution of CuSO_4 has a concentration of 2.50 M. This concentration is equivalent to

- A. 399 %m/v
- B. 39.9 %m/v
- C. 0.250 %m/v
- D. 25.0 m/v

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N Noddings



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Integrated Physics and Chemistry, Chapter 3 Activities Paradigm Accelerated Curriculum, 2005-01-01 Key topics Chemical nomenclature Lavoisiers list of elements sulfur diamonds graphite coal medieval metals platinum zinc cobalt nickel manganese molybdenum tungsten gases in the atmosphere air pressure and humidity Henry Cavendish hydrogen nitrogen fertilizers and explosives dynamite laughing gas IPC consists of twelve chapters of text and twelve companion student activity books This course introduces students to the people places and principles of physics and chemistry It is written by internationally respected scientist author John Hudson Tiner who applies the vignette approach which effectively draws readers into the text and holds attention The author and editors have deliberately avoided complex mathematical equations in order to entice students into high school level science Focus is on the people who contributed to development of the Periodic Table of the Elements Students learn to read and apply the Table while gaining insight into basic chemistry and physics This is one of our most popular courses among high school students especially those who have a history of under performance in science courses due to poor mathematical and reading comprehension skills The course is designed for two high school transcript credits Teachers may require students to complete all twelve chapters for two transcript credits or may select only six chapters to be completed for one transcript credit for Physical Science Physics or Chemistry Compliance with state and local academic essential elements should be considered when specific chapters are selected by teachers As applicable to local policies transcript credit may be assigned as follows when students complete all 12 chapters Physical Science for one credit and Chemistry for one credit or Integrated Physics and Chemistry for two credits May require supplemental local classes labs

Chemistry Unit 3 (RES) Responsive Education Solutions Staff, 2011-08-01 Chemistry students conduct field and laboratory investigations use scientific methods during investigations and make informed decisions using critical thinking and scientific problem solving Students study a variety of topics that include characteristics of matter energy transformations during physical and chemical changes atomic structure periodic table of elements behavior of gases bonding nuclear fusion and nuclear fission oxidation reduction reactions chemical equations solutes properties of solutions acids and bases and chemical reactions Students will investigate how chemistry is an integral part of our daily lives

CHEMISTRY - CHAPTER 3, 2003 **Pearson Chemistry Unit 3** Staley Wilbraham (Matta, Waterman), 2022

Chemistry Arthur A. Sandison, Southgate Publishers, Archie Gibb, Andrew Watson, Learning and Teaching Scotland, 2000

The Chemistry Book Units 3 and 4 Workbook Brett Steeples, Nicholas Stansbie, Sarah Windsor, 2019-09-18 The Chemistry Book Units 3 4 supports the development and application of key knowledge and skills for students studying science in both Queensland and greater Australia The consistent approach used in the text promotes familiarity for students supporting practice and revision of the content learnt in class and in preparation for assessment and exams The workbook revises content introduced in the classroom and provides comprehensive sets of questions to give students valuable exam

preparation and reinforce their learning

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