

Multiple Choice Answer Sheet

Module code:			
Module title:			
Date of examination:			am /pm / eve (delete as appropriate)

Delete the incorrect options for each question below, leaving only the correct option.

Question

- | | | | | | |
|-----|---|---|---|---|---|
| 1. | A | B | C | D | E |
| 2. | A | B | C | D | E |
| 3. | A | B | C | D | E |
| 4. | A | B | C | D | E |
| 5. | A | B | C | D | E |
| 6. | A | B | C | D | E |
| 7. | A | B | C | D | E |
| 8. | A | B | C | D | E |
| 9. | A | B | C | D | E |
| 10. | A | B | C | D | E |
| 11. | A | B | C | D | E |
| 12. | A | B | C | D | E |
| 13. | A | B | C | D | E |
| 14. | A | B | C | D | E |
| 15. | A | B | C | D | E |
| 16. | A | B | C | D | E |
| 17. | A | B | C | D | E |
| 18. | A | B | C | D | E |
| 19. | A | B | C | D | E |
| 20. | A | B | C | D | E |
| 21. | A | B | C | D | E |
| 22. | A | B | C | D | E |
| 23. | A | B | C | D | E |
| 24. | A | B | C | D | E |
| 25. | A | B | C | D | E |
| 26. | A | B | C | D | E |
| 27. | A | B | C | D | E |
| 28. | A | B | C | D | E |
| 29. | A | B | C | D | E |

Us History 2001 Multiple Choice Answer Key

Yan Bai



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web we used mole fractions in gen chem 1 to describe gas mixtures where the partial pressure of a gas was its mole fraction time the total pressure $p_a = x_a p_t$ mole fraction $x_a = \frac{\text{moles } a}{\text{moles total}}$ right sum $n_1 + \dots + n_n = 1$ mole fraction of substance a is often denoted by x_a and

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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web since the mole fraction of the o₂ is 0.050 it contributes 0.050 or if you will 5.0% of the total pressure thus total pressure is

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