

6. Consider 670 mL of a 4.10 M solution of $\text{Mg}(\text{NO}_3)_2$ setting in a beaker. If you evaporate all 670 mL of the solution, how many grams of solute would be left in the beaker?

407.8g

$\text{mol} = (\text{Molarity})(\text{Liters}) = (4.10\text{M})(0.670\text{L}) = 2.747 \text{ mol}$
 convert mol to g: $(2.747\text{mol})(148.3\text{g/mol}) = \mathbf{407.8\text{g}}$

Information: Molality

Molality is another way of expressing solution concentration. The symbol for molality is m.

Whereas molarity (M) represents the ratio of moles solute to liters of solution, the molality (m) is the ratio of moles solute to kilograms of solvent. It can be expressed using the following formula:

$$\text{molality} = \frac{\text{moles of solute}}{\text{kg solvent}}$$

Critical Thinking Questions

7. Consider a solution that is prepared by adding 1.34 moles of sodium nitrate to 2.5 kg of water. What is the molality of the solution?

0.536 m

$m = 1.34\text{mol} \div 2.5\text{kg} = 0.536 \text{ m}$

8. Considering the data given in question 7, is this enough data to find the molarity? If so, calculate the molarity. If not, explain why not.

No, because we don't know the total volume of the solution.

9. What is the molality of a solution that is made by dissolving 32.6 g of Na_2SO_4 in 475 g of water?

0.483 m

Convert g to mol: $32.6\text{g} \div 142.1\text{g/mol} = 0.229 \text{ mol}$
 $\text{molality} = \text{mol/kg} = 0.229\text{mol} \div 0.475\text{kg} = \mathbf{0.483 \text{ m}}$

10. Consider 2.35 moles of sodium chloride are dissolved in 1.21 kg of solution to make 1.29 liters. Calculate and compare the molarity and molality.

Molarity: 1.82 M

$M = \text{mol} \div L = 2.35\text{mol} \div 1.29 \text{ L} = 1.82 \text{ M}$

molality: 1.94 m

$m = \text{mol} \div \text{kg} = 2.35\text{mol} \div 1.21 \text{ kg} = 1.94 \text{ m}$

11. If 26.45g of Na_2SO_4 are dissolved in 1.10 kg of solution to make 1.24 L, calculate both the molarity and the molality of the resulting solution.

Convert g to mol: $26.45\text{g} \div 142.1\text{g/mol} = 0.186\text{mol}$

Molarity: 0.150 M

$M = \text{mol} \div L = 0.186\text{mol} \div 1.24 \text{ L} = 0.150 \text{ M}$

molality: 0.169 m

$m = \text{mol} \div \text{kg} = 0.186\text{mol} \div 1.10\text{kg} = 0.169 \text{ m}$

Where Can I Find Answers To Chemquest 20

Patrick Vollmar



Where Can I Find Answers To Chemquest 20:

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Table of Contents Where Can I Find Answers To Chemquest 20

1. Understanding the eBook Where Can I Find Answers To Chemquest 20
 - The Rise of Digital Reading Where Can I Find Answers To Chemquest 20
 - Advantages of eBooks Over Traditional Books
2. Identifying Where Can I Find Answers To Chemquest 20
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Where Can I Find Answers To Chemquest 20
 - Personalized Recommendations
 - Where Can I Find Answers To Chemquest 20 User Reviews and Ratings
 - Where Can I Find Answers To Chemquest 20 and Bestseller Lists

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 - Highlighting and Note-Taking Where Can I Find Answers To Chemquest 20
 - Interactive Elements Where Can I Find Answers To Chemquest 20
8. Staying Engaged with Where Can I Find Answers To Chemquest 20
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Where Can I Find Answers To Chemquest 20
9. Balancing eBooks and Physical Books Where Can I Find Answers To Chemquest 20
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Where Can I Find Answers To Chemquest 20
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Where Can I Find Answers To Chemquest 20
 - Setting Reading Goals Where Can I Find Answers To Chemquest 20
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Where Can I Find Answers To Chemquest 20
 - Fact-Checking eBook Content of Where Can I Find Answers To Chemquest 20
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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