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AP Chemistry Unit 8 Review

81 questions - 24

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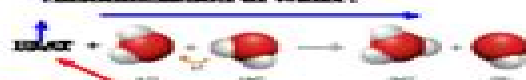
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Unit 8

Acids & Bases

8.1 Introduction to Acids & Bases

Autoionization of Water:



This is an endothermic process- increasing the heat will shift the reaction to produce more H_3O^+ and OH^- . K_w increases (more products), both pH and pOH will decrease but pure H₂O will still be neutral.

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

Only digits to the right of the decimal are significant in pH and pOH values. See

$$\text{pH} = 3.45$$

is sig. figs

In case of acidic solution $> [\text{H}^+] > [\text{OH}^-]$

$$\text{pH} < 7 \text{ at } 25^\circ\text{C}$$

In case of neutral solution $> [\text{H}^+] = [\text{OH}^-]$

$$\text{pH} = 7 \text{ at } 25^\circ\text{C}$$

In case of basic solution $> [\text{H}^+] < [\text{OH}^-]$

$$\text{pH} > 7 \text{ at } 25^\circ\text{C}$$

ION PRODUCT CONSTANT:

$$K_w = [\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ at } 25^\circ\text{C}$$

$$= K_a \times K_b$$

$$\text{pH} = -\log[\text{H}^+], \text{pOH} = -\log[\text{OH}^-]$$

$$14 = \text{pH} + \text{pOH}$$

Also known:

$$[\text{H}^+] = 10^{-\text{pH}} \text{ and } [\text{OH}^-] = 10^{-\text{pOH}}$$

3. The value of K_w for water at 0°C is 1×10^{-15} . What is the pOH of water at 0°C ?

a) 7.5

b) 8.0

c) 7.0

d) 6.5

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**Igor Krupa, Deepalekshmi
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Table of Contents V2 0 Chemistry Unit 8 Review

1. Understanding the eBook V2 0 Chemistry Unit 8 Review
 - The Rise of Digital Reading V2 0 Chemistry Unit 8 Review
 - Advantages of eBooks Over Traditional Books
2. Identifying V2 0 Chemistry Unit 8 Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an V2 0 Chemistry Unit 8 Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from V2 0 Chemistry Unit 8 Review
 - Personalized Recommendations

- V2 0 Chemistry Unit 8 Review User Reviews and Ratings
- V2 0 Chemistry Unit 8 Review and Bestseller Lists
- 5. Accessing V2 0 Chemistry Unit 8 Review Free and Paid eBooks
 - V2 0 Chemistry Unit 8 Review Public Domain eBooks
 - V2 0 Chemistry Unit 8 Review eBook Subscription Services
 - V2 0 Chemistry Unit 8 Review Budget-Friendly Options
- 6. Navigating V2 0 Chemistry Unit 8 Review eBook Formats
 - ePub, PDF, MOBI, and More
 - V2 0 Chemistry Unit 8 Review Compatibility with Devices
 - V2 0 Chemistry Unit 8 Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of V2 0 Chemistry Unit 8 Review
 - Highlighting and Note-Taking V2 0 Chemistry Unit 8 Review
 - Interactive Elements V2 0 Chemistry Unit 8 Review
- 8. Staying Engaged with V2 0 Chemistry Unit 8 Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers V2 0 Chemistry Unit 8 Review
- 9. Balancing eBooks and Physical Books V2 0 Chemistry Unit 8 Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection V2 0 Chemistry Unit 8 Review
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine V2 0 Chemistry Unit 8 Review
 - Setting Reading Goals V2 0 Chemistry Unit 8 Review
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of V2 0 Chemistry Unit 8 Review
 - Fact-Checking eBook Content of V2 0 Chemistry Unit 8 Review

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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