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The Particle Theory of Matter: More Practice

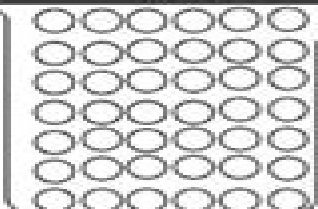
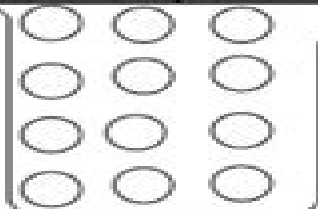
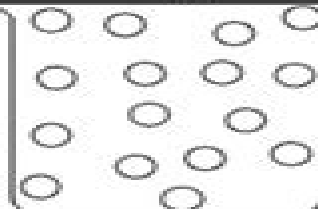
Use your prior knowledge, your notes, and your textbook (particularly section 5.1) to help you answer the following questions. You can also replay the Eureka episodes on YouTube (Molecules in Solids, Molecules in Liquids, Evaporation & Condensation, & Expansion & Contraction)

1. Define matter: Matter is anything that occupies space and has mass

2. Complete the following statements that summarize the points of the particle theory of matter:

- i. All matter is made up of extremely tiny particles
- ii. Each pure substance has its own kind of particle. These particles are identical to each other, but different from the particles of other pure substances.
- iii. Particles attract each other
- iv. Particles are always in constant motion
- v. Particles at high temperature move at a faster rate than particles at lower temperature.

3. Complete the following table to characterize particles in the three states of matter:

State	Solid	Liquid	Gas
Diagram illustrating particles			
Relative particle distance	<u>Very little space</u>	<u>bigger space</u>	<u>a lot of space</u>
Relative particle rate of motion	<u>barely moves</u>	<u>moves a little</u>	<u>moves a lot</u>
Relative particle attraction	<u>a force keeps them tight</u>	<u>a force keeps them a little tight</u>	<u>a force spreads them apart</u>
Relative volume	<u>definite volume</u>	<u>definite volume</u>	<u>no volume</u>
Shape in relation to its container	<u>no definite shape</u>	<u>no definite shape</u>	<u>no shape</u>
Compressibility	<u>non compressible</u>	<u>non compressible</u>	<u>very compressible</u>

States Of Matter And The Kinetic Theory Michigan Answers

Robert R. Redfield



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