



# Smart Car Manual 2015

**Julia Schneider**



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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of  $\text{N}_2$  what is the limiting reagent if 8.52 g are actually formed what is the

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produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have 0.216232 grams/mole 21.3 grams / 0.66 moles = 32.27 g

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web a write a balanced equation for the reaction  $Al + O_2 \rightarrow Al_2O_3$  4 Al + 3 O<sub>2</sub> → 2 Al<sub>2</sub>O<sub>3</sub> b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub> 160.0 g Al x 1 mol Al / x 2 mol Al<sub>2</sub>O<sub>3</sub> x 101.964 g Al<sub>2</sub>O<sub>3</sub> / 302.3 g Al<sub>2</sub>O<sub>3</sub> = 26.982 g Al<sub>2</sub>O<sub>3</sub> c determine the percent yield yield actual x 100 / 260.0 g x 100 = 86.01% theoretical 302.3 g

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