

**CHAPTER 10 REVIEW***States of Matter***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. Identify whether the descriptions below describe an ideal gas or a real gas.

\_\_\_\_\_ **ideal gas**      a. The gas will not condense because the molecules do not attract each other.

\_\_\_\_\_ **ideal gas**      b. Collisions between molecules are perfectly elastic.

\_\_\_\_\_ **real gas**      c. Gas particles passing close to one another exert an attraction on each other.

2. The formula for kinetic energy is
- $KE = \frac{1}{2}mv^2$
- .

- a. As long as temperature is constant, what happens to the kinetic energy of the colliding particles during an elastic collision?

**The energy is transferred between them.**

- b. If two gases have the same temperature and share the same energy but have different molecular masses, which molecules will have the greater speed?

**Those with the lower molecule mass.**

3. Use the kinetic-molecular theory to explain each of the following phenomena:

- a. A strong-smelling gas released from a container in the middle of a room is soon detected in all areas of that room.

**Gas molecules are in constant, rapid, random motion.**

- b. As a gas is heated, its rate of effusion through a small hole increases if all other factors remain constant.

**As a gas is heated, each molecule's speed increases; therefore, the molecules pass through the small hole more frequently.**

4. a. \_\_\_\_\_
- b, d, c, a**
- List the following gases in order of rate of effusion, from lowest to highest. (Assume all gases are at the same temperature and pressure.)
- 
- (a) He      (b) Xe      (c) HCl      (d) Cl
- <sub>2</sub>

# States Of Matter 13 1 Section Review Key

**Michael Hoelscher**



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