

Chemistry Worksheets Class 10 on Chapter 1 Chemical Reactions and Equations- Set 1

Q-1: Write the balanced chemical equations for each of the following reactions.

- Liquid hydrogen peroxide decomposes to form water and oxygen gas.
- Acetylene gas burns in oxygen to form carbon dioxide and water along with evolution of heat.
- When solid mercury(II) oxide is heated, liquid mercury and oxygen gas are produced.
- Marble dissolves in a hydrochloric acid to give calcium chloride, water and carbon dioxide.
- Copper sulphate on treatment with potassium iodide precipitates cuprous iodide, liberates iodine gas and also forms potassium sulphate.

Q-2: Fill in the blanks.

- _____ is the magnetic oxide of iron.
- In reactivity series, manganese lies _____ zinc.
- To the reaction between AgNO_3 solution and NaCl solution, the names given to this type of reactions are _____ and _____.
- Proteins decompose to form _____ in our body.
- If copper turnings are added to colourless AgNO_3 solution, it turns _____ after sometime.

Q-3: Which of the following displacement reactions will not take place?

- $\text{Zn(s)} + \text{FeSO}_4(\text{aq}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Fe(s)}$
- $\text{Fe(s)} + \text{CuSO}_4(\text{aq}) \rightarrow \text{FeSO}_4(\text{aq}) + \text{Cu(s)}$
- $\text{Cu(s)} + 2\text{AgNO}_3(\text{aq}) \rightarrow \text{Cu(NO}_3)_2(\text{aq}) + 2\text{Ag(s)}$
- $\text{Cu(s)} + \text{FeSO}_4(\text{aq}) \rightarrow \text{CuSO}_4(\text{aq}) + \text{Fe(s)}$

Q-4: Which of the following statements is false?

- Oxidation results in the loss of an electropositive element and the gain of an electronegative element.
- Reduction results in the loss of an electronegative element and the gain of an electropositive element.
- A species that gains electrons is referred to as a reducing agent.
- An oxidising agent is a type of specie that absorbs hydrogen atoms.

Q-5: Why are certain chemicals and medications kept in coloured bottles?

Q-6: Sakshi desired that her home be white washed. She went to the market and bought 20 kg of quick lime, which she dissolved in 60 litres of water. When she added lime to water, she noticed that the water began to boil even though it was not heated. Give an explanation for her observation. Name the product and write the corresponding chemical equation and word equation.

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LM Reder

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Table of Contents Standardized Test Practice Ch 10 Answers Chemistry

1. Understanding the eBook Standardized Test Practice Ch 10 Answers Chemistry
 - The Rise of Digital Reading Standardized Test Practice Ch 10 Answers Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Standardized Test Practice Ch 10 Answers Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Standardized Test Practice Ch 10 Answers Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Standardized Test Practice Ch 10 Answers Chemistry
 - Personalized Recommendations

- Standardized Test Practice Ch 10 Answers Chemistry User Reviews and Ratings
- Standardized Test Practice Ch 10 Answers Chemistry and Bestseller Lists
- 5. Accessing Standardized Test Practice Ch 10 Answers Chemistry Free and Paid eBooks
 - Standardized Test Practice Ch 10 Answers Chemistry Public Domain eBooks
 - Standardized Test Practice Ch 10 Answers Chemistry eBook Subscription Services
 - Standardized Test Practice Ch 10 Answers Chemistry Budget-Friendly Options
- 6. Navigating Standardized Test Practice Ch 10 Answers Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Standardized Test Practice Ch 10 Answers Chemistry Compatibility with Devices
 - Standardized Test Practice Ch 10 Answers Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Standardized Test Practice Ch 10 Answers Chemistry
 - Highlighting and Note-Taking Standardized Test Practice Ch 10 Answers Chemistry
 - Interactive Elements Standardized Test Practice Ch 10 Answers Chemistry
- 8. Staying Engaged with Standardized Test Practice Ch 10 Answers Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Standardized Test Practice Ch 10 Answers Chemistry
- 9. Balancing eBooks and Physical Books Standardized Test Practice Ch 10 Answers Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Standardized Test Practice Ch 10 Answers Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Standardized Test Practice Ch 10 Answers Chemistry
 - Setting Reading Goals Standardized Test Practice Ch 10 Answers Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Standardized Test Practice Ch 10 Answers Chemistry
 - Fact-Checking eBook Content of Standardized Test Practice Ch 10 Answers Chemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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