5.	Use the standard enthalpy of formation (ΔH_f^0) values in Appendix 1 to calculate the						
	enthalpy of reaction for the following equations. Balance the equations where						
	necessary.						
	(a)	$CH_4(g) + O_2(g) \rightarrow CO_2(g) + H_2O(\ell)$					
	(b)	$C_2H_6(g) + O_2(g) \rightarrow CO_2(g) + H_2O(\ell)$					
	(c)	$C_{12}H_{22}O_{11}$ (s, sucrose) + H_2O (ℓ) \rightarrow $C_6H_{12}O_6$ (s, glucose) + $C_6H_{12}O_6$ (s, fructose)					
	(d)	$CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$					
	(e)	HCl (aq) + NaOH (aq) \rightarrow NaCl (aq) + H ₂ O (ℓ)					
	(f)	$N_2O_5(g) \rightarrow N_2(g) + O_2(g)$					
Appei	ndix 1:	Standard	Enthalpy of Format	ion, ΔH_f^0 (kJ/mol).			
	Substance		ΔH _f (kJ/mol)	Substance	ΔH_f^0 (kJ/mol)		
CH ₄	(g)		-74.6	H ₂ O (g)	-241.8		
C ₂ H ₆	(g)		-84.0	H ₂ O (ℓ)	-285.8		

Substance	ΔH_f^0 (kJ/mol)	Substance	ΔH_f^0 (kJ/mol)
CH4 (g)	-74.6	H ₂ O (g)	-241.8
C ₂ H ₆ (g)	-84.0	H ₂ O (ℓ)	-285.8
C ₆ H ₁₂ O ₆ (s, glucose)	-1271.0	HCl (aq)	-167.2
C ₆ H ₁₂ O ₆ (s, fructose)	-1266.6	N ₂ (g)	0.0
C ₁₂ H ₂₂ O ₁₁ (s, sucrose)	-2222.1	N ₂ O ₅ (g)	+13.3
CaCO ₃ (s)	-1207.6	NaCl (aq)	-407.3
CaO (s)	-6349.0	NaOH (aq)	-44.5
CO ₂ (g)	-393.5	O2 (g)	0.0

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