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## AP CHEMISTRY FINAL EXAM REVIEW 2

1. Is energy absorbed or released in an endothermic reaction?: Absorbed.
2. What is the sign of delta H for an exothermic reaction?: Exothermic reactions release energy. The sign of delta H is negative.
3. When you touch a flask and it feels hot, is this an endothermic or exothermic process?: Exothermic; heat is being given off.
4. Define specific heat capacity.: The amount of heat required to raise 1g of a substance to 1 degrees Celsius.
5.  $q = mc(\Delta T)$  ; What does "q" stand for?: Heat (kJ)
6.  $q = mc(\Delta T)$  ; What does "c" stand for?: Specific heat capacity (J / g degrees C)
7. What is a hydrogen ion donor called?: An acid.
8. What is a hydrogen ion acceptor called?: A base.
9. Feels slippery; acid or base?: Base.
10. Tastes sour; acid or base?: Acid.
11. Tastes bitter; acid or base?: Base.
12.  $\text{pH} < 7$  ; acid, base, or neutral?: Acid
13.  $\text{pH} > 7$  ; acid, base, or neutral?: Base
14. Define neutralization.: When an acid reacts with a base to form water and a salt.  $\text{pH} = 7$ .
15. Soap; acid or base?: Base (tastes bitter)
16. Ammonia, acid or base?: Base (drain cleaners are bases).
17. Lemonade; acid or base?: Acid (tastes sour)
18. How many valence electrons are in an atom of chlorine?: 7
19. How many valence electrons does aluminum have?: 3
20. Define Reaction Rate: The decrease in concentration of reactants or the increase in concentration of products with time.
21. How does increasing temperature affect the rate of reaction?: Increasing temperature causes the kinetic energy to increase. Thus, molecules are moving faster and colliding more frequently. This causes the rate of reaction to increase.
22. If I decrease the concentration of reactant, what happens to the rate of reaction?: Decreasing concentration leads to less reactants available for collision. Therefore, reaction rate decreases.
23. How does surface area influence reaction rate?: By increasing the surface area of reactant molecules, we are increasing the space or point of contact that these molecules have available

# Spring 2013 Chemistry Final Exam Answers

**S Ashworth**



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