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AP Chemistry Unit 8 Review

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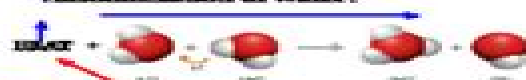
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Unit 8

Acids & Bases

8.1 Introduction to Acids & Bases

Autoionization of Water:



This is an endothermic process- increasing the heat will shift the reaction to produce more H_3O^+ and OH^- . K_w increases (more products), both pH and pOH will decrease but pure H₂O will still be neutral.

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

Only digits to the right of the decimal are significant in pH and pOH values. See

$$\text{pH} = 3.45$$

is sig. figs

In case of acidic solution $> [\text{H}^+] > [\text{OH}^-]$

$$\text{pH} < 7 \text{ at } 25^\circ\text{C}$$

In case of neutral solution $> [\text{H}^+] = [\text{OH}^-]$

$$\text{pH} = 7 \text{ at } 25^\circ\text{C}$$

In case of basic solution $> [\text{H}^+] < [\text{OH}^-]$

$$\text{pH} > 7 \text{ at } 25^\circ\text{C}$$

ION PRODUCT CONSTANT:

$$K_w = [\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ at } 25^\circ\text{C}$$

$$= K_a \times K_b$$

$$\text{pH} = -\log[\text{H}^+], \text{pOH} = -\log[\text{OH}^-]$$

$$14 = \text{pH} + \text{pOH}$$

Also known:

$$[\text{H}^+] = 10^{-\text{pH}} \text{ and } [\text{OH}^-] = 10^{-\text{pOH}}$$

3. The value of K_w for water at 0°C is 1×10^{-15} . What is the pOH of water at 0°C ?

a) 7.5

b) 8.0

c) 7.0

d) 6.5

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R J Ferrier

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