

Worksheet 13 – Intermolecular Forces

Chemical bonds are **intramolecular** forces, which hold atoms together as molecules.

Intermolecular forces exist **between** atoms and molecules, determining their **physical** properties; solubility, boiling points, melting points.

Intermolecular forces (IMF) can be ranked qualitatively using Coulomb's Law:

$$\text{force} \propto \frac{Q_1 Q_2}{r}$$

where Q_1 and Q_2 are charges and r is the distance between the charges.

The consequences are:

- the bigger the charges, the stronger the interaction
- the closer the charges, the stronger the interaction

Shown below are the structures and boiling points of nine chemical species. Use this information to answer the following questions.



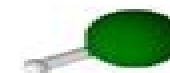
NaCl = 1413°C



N₂ = -192°C



CO₂ = -57°C



HBr = -66°C



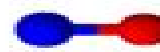
H₂O = 100°C



CH₂O = - 21°C



NH₃ = -33°C



NO = -150°C



HF = 19°C

1. Which species has the **highest** boiling point? _____
This is indicative of the **strongest IMF**.

- a. How does it differ from all of the other compounds?
- b. Look at Coulomb's law and identify each variable in the equation to the components of this species.
 Q_1 (positive charge) = _____ Q_2 (negative charge) = _____

This is called an _____ interaction (full positive and negative charges).

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Holt Rinehart & Winston

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