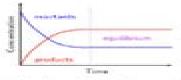
Equilibrium Practice Worksheet (Answers)

- The percent reaction of a reaction is 70%. Explain whether the equilibrium favours the reactants or products? What do you expect the magnitude of K to be? (Favours the products. Expect K value >> 0.)
- Write an equilibrium constant expression for the reactions below.
 - a. $N_{2(g)} + 3H_{2(g)} = 2NH_{3(g)} (K = (NH_3)^2 / (N_2)(H_2)^3)$
 - b. $Ba(NO_3)_{2 \text{ (aq)}} + Na_2SO_{4 \text{ (aq)}} = BaSO_{4 \text{ (s)}} + 2NaNO_{3 \text{ (aq)}}(K = (NaNO_3)^2 / (Ba(NO_3)_2)(Na_2SO_4))$
- Determine the K_c for the reaction when 0.9 M of sulfur dioxide and 0.35 M of oxygen form 1.1 M of sulfur trioxide. (K_c = 4.3)
- Using the reaction from question 3. Determine which direction the equilibrium will shift if the Q value is 1 x 10⁻⁴. (Will shift move toward the products)
- If the K_c value for an equilibrium is 3.8 x 10⁻⁶, are the products or reactants favoured? What would the K_c value for the reverse reaction? (2.6 x 10⁵)
- What would you expect the K value to be for the following graphs? (Would expect the first graph to have a K value of just above 1 and the second to have just below 1.)





Determine which direction the equilibrium will shift when the following stresses are added.

$$2SO_{3(g)} = 2SO_{2(g)} + O_{2(g)} \Delta H^{\circ} = 2198 \text{ kJ/mol}$$

- a. Temperature is increased right
- b. Pressure is decreased right
- c. Oxygen is added left
- d. A catalyst is added no shift
- e. Neon gas is added no shift
- Graph the changes when for a, c and d. (For a show increase in the products, decrease in reactants. For c should show an upward increase for O₂ and then a rate increase for reactants and rate decrease for products. For c should not see a change.)
- Which direction would you expect the equilibrium to shift if NaOH was added to the reaction below. (Would shift right)

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