

WORKSHEET ON SOLUTION CONCENTRATIONS**INTRODUCTION**

A solution is made up of two parts, the **SOLVENT** and the **SOLUTE**. The solution could be **CONCENTRATED** or **DILUTE**.

SOLVENT

The liquid which dissolves the solute. It is usually water.

SOLUTE

The substance which dissolves - often a solid, but not always.

CONCENTRATED

There is a lot of solute and a little water.

DILUTE

There is a lot of water and only a little solute.

When we measure out 50cm^3 of a solution, e.g., hydrochloric acid, we measure out some solvent (water) and some acid together. In any reactions of the acid, it is the HCl particles which are involved. The water is just a carrier for the acid, so when we measure out a volume of the solution, we want to know how much acid it contains. To do this, we need to define the **CONCENTRATION** of the solution. This tells us *how much* solute there is in a *fixed volume* of the solution. In chemistry, "*how much*" means **MOLES** and the "*fixed volume*" is taken as 1 cubic decimetre (1dm^3).

[NB: 1dm^3 is 1000cm^3 and is the same volume as 1 litre.]

DEFINITION

Concentration of solution (moles per dm^3)	=	<u>number of moles of solute</u> <u>volume of solution in dm^3</u>
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- You must learn this relationship and be able to use it and rearrange it to calculate moles of solute or volume of solution, given the other two quantities.
- Remember that the volume must be in dm^3 , so convert from cm^3 if necessary.
- The units for concentration are *moles per cubic decimetre* which is written mol.dm^{-3}
- Lab bottles use the letter *M* (meaning *molar*) and this is the same thing as mol.dm^{-3}

RELATIONSHIPS

Use the triangle opposite to help you answer the questions which follow. Some worked examples are given to help you.



Examples:-

- 1) What is the concentration of a solution which contains 2 moles of copper (II) sulphate in 4dm^3 of solution?

Ans: concentration = moles ÷ volume = $2 \div 4 = \underline{0.5 \text{ mol.dm}^{-3}}$

- 2) How many moles of sodium hydroxide are there in 200cm^3 of a solution of concentration 0.2 mol.dm^{-3} ? (Remember that $200\text{cm}^3 = 200 \div 1000 = 0.2\text{dm}^3$)

Ans: moles = concentration × volume = $0.2 \times 0.2 = \underline{0.04 \text{ moles}}$

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Stephen Lamb, Russell W. Rumberger



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Table of Contents Skill Practice 39 Concentration Practice Answers

1. Understanding the eBook Skill Practice 39 Concentration Practice Answers
 - The Rise of Digital Reading Skill Practice 39 Concentration Practice Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Skill Practice 39 Concentration Practice Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Practice 39 Concentration Practice Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Practice 39 Concentration Practice Answers

- Personalized Recommendations
 - Skill Practice 39 Concentration Practice Answers User Reviews and Ratings
 - Skill Practice 39 Concentration Practice Answers and Bestseller Lists
5. Accessing Skill Practice 39 Concentration Practice Answers Free and Paid eBooks
 - Skill Practice 39 Concentration Practice Answers Public Domain eBooks
 - Skill Practice 39 Concentration Practice Answers eBook Subscription Services
 - Skill Practice 39 Concentration Practice Answers Budget-Friendly Options
 6. Navigating Skill Practice 39 Concentration Practice Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Skill Practice 39 Concentration Practice Answers Compatibility with Devices
 - Skill Practice 39 Concentration Practice Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Skill Practice 39 Concentration Practice Answers
 - Highlighting and Note-Taking Skill Practice 39 Concentration Practice Answers
 - Interactive Elements Skill Practice 39 Concentration Practice Answers
 8. Staying Engaged with Skill Practice 39 Concentration Practice Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skill Practice 39 Concentration Practice Answers
 9. Balancing eBooks and Physical Books Skill Practice 39 Concentration Practice Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Practice 39 Concentration Practice Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Skill Practice 39 Concentration Practice Answers
 - Setting Reading Goals Skill Practice 39 Concentration Practice Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Skill Practice 39 Concentration Practice Answers

- Fact-Checking eBook Content of Skill Practice 39 Concentration Practice Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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