

Moles to Particles Conversions

Name _____

Complete the following conversions. For FULL CREDIT, you must show ALL of your work!

1. Determine the number of atoms in 0.58 moles of Se.

$$0.58 \text{ mol Se} \times \frac{6.02 \times 10^{23} \text{ atoms Se}}{1 \text{ mol Se}} = 3.5 \times 10^{23} \text{ atoms Se}$$

2. How many moles are in 6.80×10^{24} formula units $\text{Ba}(\text{NO}_3)_2$?

$$6.80 \times 10^{24} \text{ formula units Ba(NO}_3)_2 \times \frac{1 \text{ mol Ba(NO}_3)_2}{6.02 \times 10^{23} \text{ formula units Ba(NO}_3)_2} = 11.3 \text{ mol Ba(NO}_3)_2$$

3. Determine the number of molecules in 1.25 moles of O_2 gas.

$$1.25 \text{ mol O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules O}_2}{1 \text{ mol O}_2} = 7.52 \times 10^{23} \text{ molecules O}_2$$

4. How many moles are in 5.38×10^{24} formula units MgBr_2 ?

$$5.38 \times 10^{24} \text{ formula units MgBr}_2 \times \frac{1 \text{ mol MgBr}_2}{6.02 \times 10^{23} \text{ formula units MgBr}_2} = 8.94 \text{ mol MgBr}_2$$

5. Determine the number of formula units in 0.688 moles of AgNO_3 .

$$0.688 \text{ mol AgNO}_3 \times \frac{6.02 \times 10^{23} \text{ formula units AgNO}_3}{1 \text{ mol AgNO}_3} = 4.14 \times 10^{23} \text{ formula units AgNO}_3$$

6. Convert 8.46×10^{24} molecules of ethane (C_2H_6) to moles.

$$8.46 \times 10^{24} \text{ molecules C}_2\text{H}_6 \times \frac{1 \text{ mol C}_2\text{H}_6}{6.02 \times 10^{23} \text{ molecules C}_2\text{H}_6} = 14.1 \text{ mol C}_2\text{H}_6$$

7. Determine the number of formula units in 1.48 moles of NaF .

$$1.48 \text{ mol NaF} \times \frac{6.02 \times 10^{23} \text{ formula units NaF}}{1 \text{ mol NaF}} = 8.91 \times 10^{23} \text{ formula units NaF}$$

8. Convert 8.2×10^{23} molecules of water to moles.

$$8.2 \times 10^{23} \text{ molecules H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{6.02 \times 10^{23} \text{ molecules H}_2\text{O}} = 1.4 \text{ mol H}_2\text{O}$$

9. How many formula units are in 15 moles of Li_2SO_4 ?

$$15 \text{ mol Li}_2\text{SO}_4 \times \frac{6.02 \times 10^{23} \text{ formula units Li}_2\text{SO}_4}{1 \text{ mol Li}_2\text{SO}_4} = 9.0 \times 10^{24} \text{ formula units Li}_2\text{SO}_4$$

10. Calculate the number of moles in 7.2×10^{23} atoms of sulfur.

$$7.2 \times 10^{23} \text{ atoms S} \times \frac{1 \text{ mol S}}{6.02 \times 10^{23} \text{ atoms S}} = 0.12 \text{ mol S}$$

*Answer rounded using significant figures even/odd rounding rules.

Skills Practice 31 Moles And Formulas

Ralph Thomas Myers



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Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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