

Chapter 12 Review 2

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- 1 A mutual electrical attraction between the nuclei and valence electrons of different atoms that binds the atoms together is called a(n)
 - a. dipole.
 - b. Lewis structure.
 - c. chemical bond.
 - d. London force.
- 2 The electrons involved in the formation of a chemical bond are called
 - a. dipoles.
 - b. *s* electrons.
 - c. Lewis electrons.
 - d. valence electrons.
- 3 A chemical bond resulting from the electrostatic attraction between positive and negative ions is called a(n)
 - a. covalent bond.
 - b. ionic bond.
 - c. charged bond.
 - d. dipole bond.
- 4 If two covalently bonded atoms are identical, the bond is
 - a. nonpolar covalent.
 - b. polar covalent.
 - c. nonionic.
 - d. coordinate covalent.
- 5 If the atoms that share electrons have an unequal attraction for the electrons, the bond is called
 - a. nonpolar.
 - b. polar.
 - c. ionic.
 - d. dipolar.
- 6 A covalent bond results when _____ are shared.
 - a. ions
 - b. Lewis structures
 - c. electrons
 - d. dipoles
- 7 Nonpolar covalent bonds are not common because
 - a. one atom usually attracts electrons more strongly than the other.
 - b. ions always form when atoms join.
 - c. the electrons usually remain equally distant from both atoms.
 - d. dipoles are rare in nature.
- 8 The greater the electronegativity difference between two bonded atoms, the greater the percentage of
 - a. ionic character.
 - b. covalent character.
 - c. metallic character.
 - d. electron sharing.
- 9 The pair of elements that forms a bond with the least ionic character is
 - a. Na and Cl.
 - b. H and Cl.
 - c. O and Cl.
 - d. Br and Cl.
10. In which of these compounds is the bond between the atoms NOT a nonpolar covalent bond?
 - a. Cl_2
 - b. H_2
 - c. HCl
 - d. O_2

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G Psacharopoulos



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