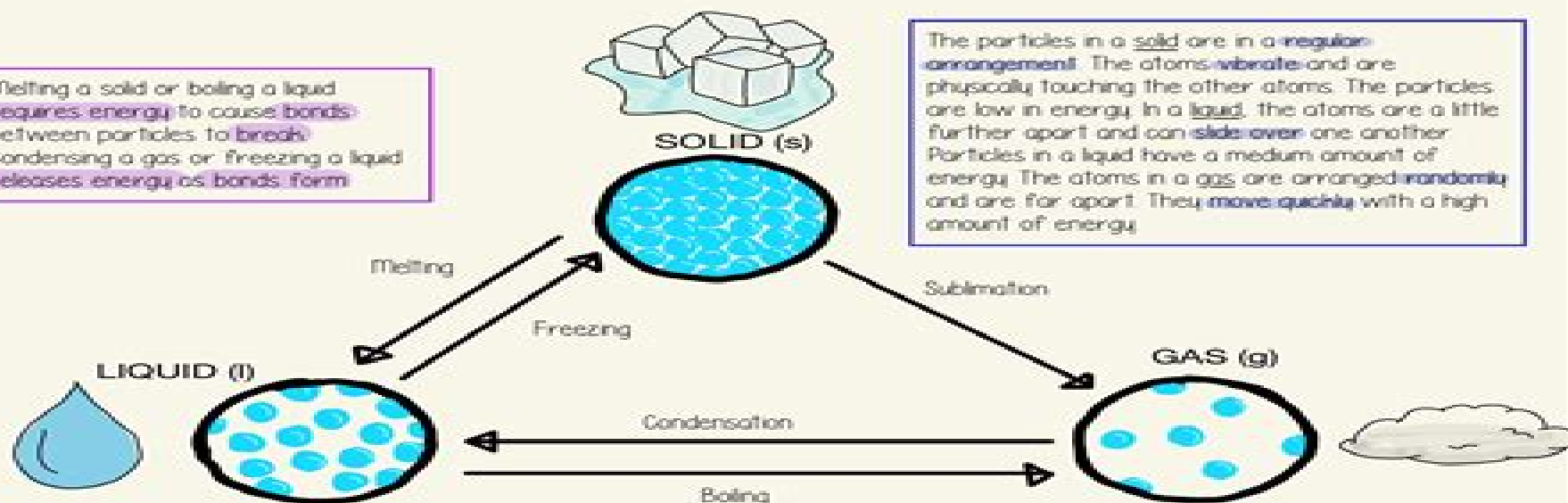
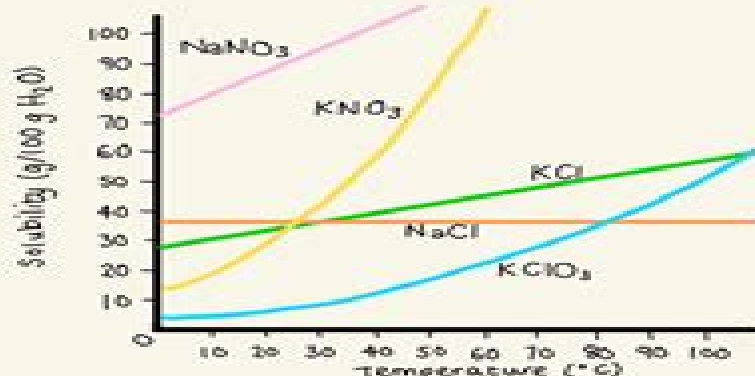


STATES OF MATTER

Melting a solid or boiling a liquid requires energy to cause bonds between particles to break. Condensing a gas or freezing a liquid releases energy as bonds form.



Solutions are formed when a solute (a solid or gas) are dissolved in a solvent (a liquid). When the solvent has dissolved as much solute as possible, it is saturated. The ability of a solute to dissolve in a solvent improves with increasing temperature. The solubility of a substance at different temperatures can be plotted on a solubility curve. Solutions are described as aqueous and are given the state symbol (aq).



Questions

1. Describe the arrangement of particles in a solid.
2. In which state do particles have the greatest energy?
3. What name is given to the state change when liquid water is converted into water vapour?
4. Look at the solubility curve on the left. Which solute has the greatest solubility at room temperature (20 degrees)?
5. Which state symbol is used most often for the elements of the Periodic Table at room temperature?
6. Write an equation, including state symbols, for the change of state from steam to water.

A solubility curve showing how soluble different solutes are with increasing temperature.

State Of Matter In Chemistry

Chris Oxlade



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