

Use the standard enthalpy of formation (ΔH_f°) values in **Appendix 1** to calculate the enthalpy of reaction for the following equations. Balance the equations where necessary.

- 5.
- (a) $\text{CH}_4 (\text{g}) + \text{O}_2 (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + \text{H}_2\text{O} (\text{l})$
 - (b) $\text{C}_2\text{H}_6 (\text{g}) + \text{O}_2 (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + \text{H}_2\text{O} (\text{l})$
 - (c) $\text{C}_{12}\text{H}_{22}\text{O}_{11} (\text{s, sucrose}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 (\text{s, glucose}) + \text{C}_6\text{H}_{12}\text{O}_6 (\text{s, fructose})$
 - (d) $\text{CaCO}_3 (\text{s}) \rightarrow \text{CaO} (\text{s}) + \text{CO}_2 (\text{g})$
 - (e) $\text{HCl} (\text{aq}) + \text{NaOH} (\text{aq}) \rightarrow \text{NaCl} (\text{aq}) + \text{H}_2\text{O} (\text{l})$
 - (f) $\text{N}_2\text{O}_5 (\text{g}) \rightarrow \text{N}_2 (\text{g}) + \text{O}_2 (\text{g})$

Appendix 1: Standard Enthalpy of Formation, ΔH_f° (kJ/mol).

Substance	ΔH_f° (kJ/mol)
$\text{CH}_4 (\text{g})$	-74.6
$\text{C}_2\text{H}_6 (\text{g})$	-84.0
$\text{C}_6\text{H}_{12}\text{O}_6 (\text{s, glucose})$	-1271.0
$\text{C}_6\text{H}_{12}\text{O}_6 (\text{s, fructose})$	-1266.6
$\text{C}_{12}\text{H}_{22}\text{O}_{11} (\text{s, sucrose})$	-2222.1
$\text{CaCO}_3 (\text{s})$	-1207.6
$\text{CaO} (\text{s})$	-6349.0
$\text{CO}_2 (\text{g})$	-393.5

Substance	ΔH_f° (kJ/mol)
$\text{H}_2\text{O} (\text{g})$	-241.8
$\text{H}_2\text{O} (\text{l})$	-285.8
$\text{HCl} (\text{aq})$	-167.2
$\text{N}_2 (\text{g})$	0.0
$\text{N}_2\text{O}_5 (\text{g})$	+13.3
$\text{NaCl} (\text{aq})$	-407.3
$\text{NaOH} (\text{aq})$	-44.5
$\text{O}_2 (\text{g})$	0.0

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