

Gas Stoichiometry Quiz SCH3U

Name: _____

 T C A **Assessment of Learning (Product)**
 5 4 12

$$R = \frac{8.31 \text{ kPa} \cdot \text{L}}{\text{mol} \cdot \text{K}}$$

$$1 \text{ atm} = 101.325 \text{ kPa} = 760 \text{ torr} = 760 \text{ mmHg}$$

$$\text{SATP} = 25^\circ\text{C and } 100 \text{ kPa}$$

$$\text{STP} = 0^\circ\text{C and } 101.3 \text{ kPa}$$

1. What pressure would result if 45 grams of chlorine gas was heated to 70°C in a 50 mL container? (3T, 1C)

2. Wood is composed mainly of cellulose, $\text{C}_6\text{H}_{10}\text{O}_5$. 1 kg of wood is burned with 30 L of oxygen in a campfire to produce carbon dioxide gas and water vapour. Write the balanced chemical equation. (2T)



a. If the temperature of the fire was 300°C and the atmospheric pressure at the time was 1 atm calculate the limiting reactant? (3A)

b. What volume of H_2O would be produced at 300°C and 1 atm? (3A, 1C)

c. What volume of CO_2 would be produced at STP conditions? (3A, 1C)

Stoichiometry Quiz Unit 8 Quiz B

**William Setchel Learned, Ben DeKalbe
Wood**



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