

Worksheet #3—Stoichiometry

Directions: Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation: $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

$$\frac{14.8 \text{ moles Al}}{4 \text{ mole Al}} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole Al}} = \underline{11.1 \text{ mol O}_2}$$

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride (ZnCl_2) are produced?



$$\frac{20 \text{ g Zn}}{65 \text{ g Zn}} \times \frac{1 \text{ mol Zn}}{1 \text{ mole Zn}} \times \frac{1 \text{ mole ZnCl}_2}{1 \text{ mole Zn}} \times \frac{136 \text{ g ZnCl}_2}{1 \text{ mol ZnCl}_2}$$

$$= \underline{41.8 \text{ g ZnCl}_2}$$

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?



$$\frac{4.0 \text{ g AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mol AgNO}_3}{1 \text{ mole AgNO}_3} \times \frac{1 \text{ mole Cu}}{2 \text{ mol AgNO}_3} \times \frac{63.5 \text{ g Cu}}{1 \text{ mole Cu}}$$

2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO_3).



$$\frac{0.61 \text{ moles KClO}_3}{2 \text{ mole KClO}_3} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole KClO}_3} =$$

$$\underline{0.915 \text{ mol O}_2}$$

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed



$$\frac{10 \text{ g NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole Cl}_2}{2 \text{ mol NaCl}} \times \frac{71 \text{ g Cl}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{6.07 \text{ g Cl}_2}$$

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?



$$\frac{15.1 \text{ g Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mol Cl}_2}{1 \text{ mole Cl}_2} \times \frac{1 \text{ mole H}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{0.21 \text{ moles H}_2}$$

Stoichiometry Practice Problem Sheet 2 Answers

Sigurd Skogestad



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Table of Contents Stoichiometry Practice Problem Sheet 2 Answers

1. Understanding the eBook Stoichiometry Practice Problem Sheet 2 Answers
 - The Rise of Digital Reading Stoichiometry Practice Problem Sheet 2 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Practice Problem Sheet 2 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Practice Problem Sheet 2 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Practice Problem Sheet 2 Answers
 - Personalized Recommendations
 - Stoichiometry Practice Problem Sheet 2 Answers User Reviews and Ratings
 - Stoichiometry Practice Problem Sheet 2 Answers and Bestseller Lists
5. Accessing Stoichiometry Practice Problem Sheet 2 Answers Free and Paid eBooks
 - Stoichiometry Practice Problem Sheet 2 Answers Public Domain eBooks
 - Stoichiometry Practice Problem Sheet 2 Answers eBook Subscription Services
 - Stoichiometry Practice Problem Sheet 2 Answers Budget-Friendly Options
6. Navigating Stoichiometry Practice Problem Sheet 2 Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Stoichiometry Practice Problem Sheet 2 Answers Compatibility with Devices
 - Stoichiometry Practice Problem Sheet 2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Practice Problem Sheet 2 Answers
 - Highlighting and Note-Taking Stoichiometry Practice Problem Sheet 2 Answers
 - Interactive Elements Stoichiometry Practice Problem Sheet 2 Answers
 8. Staying Engaged with Stoichiometry Practice Problem Sheet 2 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Practice Problem Sheet 2 Answers
 9. Balancing eBooks and Physical Books Stoichiometry Practice Problem Sheet 2 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Practice Problem Sheet 2 Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Stoichiometry Practice Problem Sheet 2 Answers
 - Setting Reading Goals Stoichiometry Practice Problem Sheet 2 Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Stoichiometry Practice Problem Sheet 2 Answers
 - Fact-Checking eBook Content of Stoichiometry Practice Problem Sheet 2 Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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