

Stoichiometry Worksheet #1

1. Aluminum chloride, AlCl_3 , is used as a catalyst in various industrial reactions. It is prepared from hydrogen chloride gas and aluminum metal shavings.



1A. How many moles of AlCl_3 can be prepared from 3.5 moles of hydrogen chloride gas with an excess of aluminum?

Ans: **1.2 mol AlCl_3**

1B. How many moles of hydrogen gas would be produced from the use of 8.5 moles of aluminum with an excess of hydrogen chloride?

Ans: **13 mol H_2**

2. When dinitrogen pentoxide, N_2O_5 , a white solid, is heated, it decomposes to nitrogen dioxide and oxygen.



2A. How many moles of nitrogen dioxide can be formed from the decomposition of 1.25 g of N_2O_5 ?

Ans: **0.0231 mol NO_2**

2B. How many grams of oxygen can be formed from the decomposition of 2.3 g of N_2O_5 ?

Ans: **0.34 g O_2**

3. Chlorine is prepared from sodium chloride by electrochemical decomposition. Formerly chlorine was produced by heating hydrochloric acid with pyrolusite [manganese dioxide or manganese(IV) oxide, MnO_2], a common manganese ore. Small amounts of chlorine may be prepared in the laboratory by the same reaction.



3A. How many grams of HCl react with 5.00 g of MnO_2 , according to the equation?

Ans: **8.39 g HCl**

3B. If a chemist wanted to prepare 100. g of chlorine, how many grams of MnO_2 are needed, assuming there is more than enough hydrochloric acid?

Ans: **123 g MnO_2**

3C. How many molecules of water are produced from the reaction of 5.0 g of HCl ?

Ans: **4.1×10^{23} molec. H_2O**

4. Sodium is a soft, reactive metal that instantly reacts with water to give hydrogen gas and a solution of sodium hydroxide, NaOH .

4A. How many grams of sodium metal are needed to give 7.81 g of hydrogen by this reaction?

Ans: **178 g Na**

4B. How many sodium atoms are needed to react with 1.25×10^{24} molecules of water?

Ans: **1.25×10^{24} atoms Na**

Stoichiometry Mixed Problems Worksheet 1

S Ashworth



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