Skill - Student can use a balanced equation to determine mole to mole ratios.

- 1) Given this equation, N. + 3 H. + 2 NH., write the following molar ratios.

 - ber Marc Well, Both
 - CONTRACT MARKS THE PARTY
- Given the following equation: 8 H₂ + S₂ → 8 H₂S, write the following molar ratios:
 - A H. H. H. C. C. C.
 - by Mary Co. But II.
 - CONTRACTOR OF THE REAL PROPERTY.
- 30 Answer the following questions for this equation: 2 H. + O. + 2 Hub
 - What is the H_a: H_aO molar ratio?
 - b. With 20 moles of H_a on hand & plenty of O_a, how many moles of H_aO could you make?
 - c. What is the Out Mub molar ratio?
 - d. With 29 moles of 0₃ and enough H_a, how many moles of H₂O could you make?
- 40 American the following questions for this equation: N. + 3 M. + 2 NH.
 - a. If you used 1 mole of N_s, how many makes of NH_s could be produced?
 - if 50 moles of fift, were produced, how many moles of N, would be required?
 10mol 50H, v (1)mol N, (2mol 50H,) = 5 mol N.
 - c. If 3.00 moles of Humere used, how many moles of NH, would be made?
 - d. if 0.600 moles of NH, were produced, how many moles of M, are required? 0.600 mol NH, x (Simol H,/2 mol NH,) = 0.9 mol H,

Skill - Student can use a balanced equation, given the grams of reactant, to calculate the orange of another substance

5) Magnesium reacts with hydrochloric acid according to the following equation:

He fet + 2 HCl fact + He/I, tack + H, foll

72.92 grams of hydrochiloric acid react with excess magnesium, how many moles of hydrogen gas will be produced?

72.92g HCl x 1mml HCl x 1.mml H, = 1.000weel H, 35.46gHCl 2mml HCl

- 6) Aluminum reacts with hydrobromic acid to produce aluminum bromide and hydrogen gas. Write a balanced equation for the reaction and calculate the number of moles of hydrobromic acid required to react with 23.47 grams of aluminum.
 2 Alia) + 6 Http://eq.) 2 Althylio) + 3 H.(q)
 - Still King Freed

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