

Stoichiometry: Assessment Review

Name _____

Skill – Student can use a balanced equation to determine mole to mole ratios

- 1) Given this equation: $N_2 + 3 H_2 \rightarrow 2 NH_3$, write the following molar ratios:
- $N_2 : H_2$ **1:3**
 - $N_2 : NH_3$ **1:2**
 - $H_2 : NH_3$ **3:2**
- 2) Given the following equation: $8 H_2 + S_8 \rightarrow 8 H_2S$, write the following molar ratios:
- $H_2 : H_2S$ **8:8**
 - $H_2 : S_8$ **8:1**
 - $H_2S : S_8$ **8:1**
- 3) Answer the following questions for this equation: $2 H_2 + O_2 \rightarrow 2 H_2O$
- What is the $H_2 : H_2O$ molar ratio?
2:2
 - With 20 moles of H_2 on hand & plenty of O_2 , how many moles of H_2O could you make?
20 moles
 - What is the $O_2 : H_2O$ molar ratio?
1:2
 - With 20 moles of O_2 and enough H_2 , how many moles of H_2O could you make?
40 moles
- 4) Answer the following questions for this equation: $N_2 + 3 H_2 \rightarrow 2 NH_3$
- If you used 1 mole of N_2 , how many moles of NH_3 could be produced?
2moles
 - If 20 moles of NH_3 were produced, how many moles of N_2 would be required?
10moles $NH_3 \times (1\text{mole } N_2 / 2\text{moles } NH_3) = 5\text{ mole } N_2$
 - If 3.00 moles of H_2 were used, how many moles of NH_3 would be made?
2moles
 - If 0.600 moles of NH_3 were produced, how many moles of H_2 are required?
0.600moles $NH_3 \times (3\text{moles } H_2 / 2\text{moles } NH_3) = 0.9moles H_2$

Skill – Student can use a balanced equation, given the grams of reactant, to calculate the grams of another substance

- 5) Magnesium reacts with hydrochloric acid according to the following equation:



If 72.92 grams of hydrochloric acid react with excess magnesium, how many moles of hydrogen gas will be produced?

$$72.92g \text{ HCl} \times \frac{1\text{mol HCl}}{35.46g \text{ HCl}} \times \frac{1\text{mol } H_2}{2\text{mol HCl}} = \mathbf{1.002moles H_2}$$

- 6) Aluminum reacts with hydrobromic acid to produce aluminum bromide and hydrogen gas. Write a balanced equation for the reaction and calculate the number of moles of hydrobromic acid required to react with 25.47 grams of aluminum.



$$25.47g \text{ Al} \times \frac{1\text{mol Al}}{26.98g \text{ Al}} \times \frac{6\text{mol HBr}}{2\text{mol Al}} = \mathbf{2.810moles HBr}$$

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