

Stoichiometry Lab: Vinegar and Baking Soda

Do Not Write On This Sheet

Purpose: To predict the amount of Carbon Dioxide gas that should be produced in a chemical reaction; then calculate the % yield.



Materials: Baking Soda (NaHCO_3), Vinegar (CH_3COOH), and 2 plastic cups, scale.

Procedure: Find and record the mass of cup A. With cup A still on the scale, add approximately 10.0 g of baking soda to the cup. **The mass does not have to be exact, as long as you carefully record your results.**

Place cup B on the scale, weigh and record approximately 50.0 g of vinegar. **The mass does not have to be exact, as long as you carefully record your results.**

Slowly add vinegar to cup A until the reaction has stopped. **DO NOT** add all of the vinegar, just enough to complete the reaction. Reweigh and record both cup A and B. Calculate the mass of CO_2 that escaped.

Data:

- a. Mass of Cup A _____ g
- b. Mass of Cup A and baking soda _____ g
- c. Calculate mass of baking soda (b - a) _____ g
- d. Mass of Cup B with vinegar _____ g
- e. Mass of Cup B after reaction _____ g
- f. Calculate mass of vinegar poured into Cup A (d - e) _____ g
- g. Mass of Cup A after reaction _____ g
- h. Calculate mass of product after reaction (g - a) _____ g
- i. Calculate baking soda + vinegar (c + f) _____ g
- j. Calculate mass of CO_2 lost (i - h) _____ g

Stoichiometry Problems For Vinegar And Baking Soda

Ralph Thomas Myers



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eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

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