

STOICHIOMETRY: MASS-MASS PROBLEMS

Name _____



How many grams of potassium chloride are produced if 25 g of potassium chlorate decompose?

$$25\text{g KClO}_3 \left(\frac{1\text{ mol}}{122.5\text{g}} \right) \left(\frac{2\text{ mol KCl}}{2\text{ mol KClO}_3} \right) \left(\frac{74.5\text{g KCl}}{1\text{ mol KCl}} \right) = 13.2\text{g KCl}$$



How many grams of hydrogen are necessary to react completely with 50.0 g of nitrogen in the above reaction?

$$50\text{g N}_2 \left(\frac{1\text{ mol N}_2}{28\text{g N}_2} \right) \left(\frac{3\text{ mol H}_2}{1\text{ mol N}_2} \right) \left(\frac{2\text{g H}_2}{1\text{ mol H}_2} \right) = 10.7\text{g H}_2$$

3. How many grams of ammonia are produced in the reaction in Problem 2?

$$50\text{g N}_2 \quad 50\text{g} + 10.7 = \boxed{60.7\text{g NH}_3}$$



How many grams of silver chloride are produced from 5.0 g of silver nitrate reacting with an excess of barium chloride?

$$5.0\text{g AgNO}_3 \left(\frac{1\text{ mol}}{170\text{g}} \right) \left(\frac{1\text{ mol AgCl}}{1\text{ mol AgNO}_3} \right) \left(\frac{143\text{g}}{1\text{ mol AgCl}} \right) = 4.21\text{g AgCl}$$

5. How much barium chloride is necessary to react with the silver nitrate in Problem 4?

$$5.0\text{g AgNO}_3 \left(\frac{1\text{ mol}}{170\text{g AgNO}_3} \right) \left(\frac{1\text{ mol BaCl}_2}{2\text{ mol AgNO}_3} \right) \left(\frac{207\text{g}}{1\text{ mol BaCl}_2} \right) = \boxed{5.0\text{g BaCl}_2}$$

9-7

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Sunil Tanna



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Bite-Sized Chemistry Calculations John Lambert, 2024-07-10 Bite Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry focusing on a few topics at a time to facilitate comprehension. Written by an experienced chemistry educator each book in the series has been tailored to fully meet the needs of students at all levels especially those taking college level general chemistry courses as well as those following various O level curricula worldwide. This part of the series explores the different types of problems and calculations encountered in mass the mole and stoichiometry including the determination of formulae of ionic compounds relative formula masses mass and percent compositions of compounds all aspects of mole calculations empirical and molecular formulae calculations based on chemical equations limiting reagents gas stoichiometry and percent yield. The series is packed with many salient features that are meant to facilitate both teaching and learning. Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices. With this book you will be well prepared for your exams and boost your performance.

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Answers to Practice Problems
Appendices

You Can Do Chemistry Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers. Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more. Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018.

Solving Problems in Chemistry Rod O'Connor, Charles Mickey, Alton Hassell, 1977

Chemistry Problems David E. Newton, 2001 This edition includes acid base chemistry and thermochemistry. Chemistry Problems is the authoritative resource for practice problems covering all the essentials. Includes Atomic structure Stoichiometry Solutions chemistry and Electrochemistry. Literally thousands of problems in this compendium build proficiency analytical skills and math skills. The text includes a complete answer key and reference to applicable web sites.

A Stoichiometry Unit David Callaghan, 2004

Stoichiometry Problems, 2016 You extend your study of stoichiometry to consider more complex problems involving volume molecules and energy.

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Chemical Problem-solving by Dimensional Analysis Arnold B. Loebel, 1987

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Calculating in Chemistry Cosmas C

Mbuzi,2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance This is particularly useful to a beginner Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry

Stoichiometry and Structure. Freshman Chemistry Problems and how to Solve Them Michell Joseph Sienko,1964 *Problems and Worked Examples in Chemistry* Albert Holderness,John Lambert,1978 **Chemistry,**

Problem-Solving Worktext James E. Brady,Frederick Senese,2002-06-13 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

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