

EXPERIMENT 2 PRE-LAB QUESTIONS (20 points)

Stoichiometry and Thermochemistry

1. There are three possible unknowns for this experiment. Write balanced chemical equations for the reaction of each of these unknowns with sodium hypochlorite. (3 points)



2. If the stoichiometric ratio of your reaction were 1:1, what volumes of hypochlorite and your unknown would you use to get the highest temperature possible? What if it were 3:1? 4:1? (Hint: Remember the total volume will be 10 mL.) (3 points)

1:1 = 2 moles $\text{NaOCl} = \frac{1}{2} \text{ mole} = 0.5 \times 10 \text{ ml} = 5 \text{ ml}$
Unknown = $\frac{1}{2} \text{ mole} = 0.5 \times 10 \text{ ml} = 5 \text{ ml}$

3:1, total of 4 moles $\text{NaOCl} = \frac{3}{4} \text{ mole} = \frac{3}{4} \times 10 \text{ ml} = 7.5 \text{ ml}$
Unknown = 2.5 ml

4:1 total moles = 5, $\text{NaOCl} = \frac{4}{5} \rightarrow 8 \text{ ml}$
Unknown = 2 ml

3. If the stoichiometric ratio of your reaction were 1:1, what volumes would you use to collect four data points on each side of the highest temperature (stoichiometric ratio)? What if it were 3:1? 4:1? (3 points)

10, 20, 30, 40 ml

4. Finally, if you didn't know what the stoichiometric ratio of hypochlorite and your unknown was, what would be your approach? What volumes of hypochlorite and unknown would you start with? What other volumes would you use? How would you know you had enough data? (2 points)

You'd have to try multiple volumes because you don't know what number of moles they are reacting. If you try it with 1, first, you get some product, and if any reactant is left that indicates a mismatch in the number of moles reacting, and then you can try different volumes to check whether the finding is correct.

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L Reisser



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