Stoichiometry Worksheet # 1

 Glucose is used as a source of energy by the human body. The overall reaction in the body is C6H12O6+6O2 → 6CO2 + 6H2O

Calculate the number of moles of oxygen needed to oxidize 12.5 g of C6H12O6.

Ammonia is synthesized from hydrogen and nitrogen according to the following equation.

$$N_2 + 3H_2 \rightarrow 2NH_3$$

If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH₃) can be produced?

3) Assume that in the decomposition of potassium chlorate, KCLO₃,

80.5 moles of O: form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,

9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" PtCl₂(NH₂)₂, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction

$$K_2PtCl_4 + 2NH_3 \longrightarrow 2KCl + PtCl_2(NH_3)_2$$

- A. How many moles of "cisplatin" can be produced from 2.50 g K2PtCl4?
- B. How many moles of NH3 would be needed if 3.8 moles of K2PtCl4 were present?

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Table of Contents Stoichiometry Chapter 9 Answers

- 1. Understanding the eBook Stoichiometry Chapter 9 Answers
 - The Rise of Digital Reading Stoichiometry Chapter 9 Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Stoichiometry Chapter 9 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Chapter 9 Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Stoichiometry Chapter 9 Answers
 - Personalized Recommendations
 - Stoichiometry Chapter 9 Answers User Reviews and Ratings
 - Stoichiometry Chapter 9 Answers and Bestseller Lists

- 5. Accessing Stoichiometry Chapter 9 Answers Free and Paid eBooks
 - Stoichiometry Chapter 9 Answers Public Domain eBooks
 - Stoichiometry Chapter 9 Answers eBook Subscription Services
 - Stoichiometry Chapter 9 Answers Budget-Friendly Options
- 6. Navigating Stoichiometry Chapter 9 Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Stoichiometry Chapter 9 Answers Compatibility with Devices
 - Stoichiometry Chapter 9 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Chapter 9 Answers
 - Highlighting and Note-Taking Stoichiometry Chapter 9 Answers
 - Interactive Elements Stoichiometry Chapter 9 Answers
- 8. Staying Engaged with Stoichiometry Chapter 9 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Chapter 9 Answers
- 9. Balancing eBooks and Physical Books Stoichiometry Chapter 9 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Chapter 9 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry Chapter 9 Answers
 - Setting Reading Goals Stoichiometry Chapter 9 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry Chapter 9 Answers
 - Fact-Checking eBook Content of Stoichiometry Chapter 9 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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