

# Stoichiometry Balloon Races

## Limiting and Excess Reactants



### Introduction

If a little is good, more is better, right? Increasing the mass of a reactant in a chemical reaction may not increase the amount of product that can be formed—the yield may be “limited” based on the mole ratio of reactants. Demonstrate the concept of limiting and excess reactants by comparing the amount of carbon dioxide obtained when varying amounts of sodium bicarbonate react with a constant amount of acetic acid. It’s off to the races . . .

### Concepts

- Stoichiometry
- Limiting reactant
- Mole ratio
- Excess reactant

### Materials

|                                                             |                           |
|-------------------------------------------------------------|---------------------------|
| Acetic acid solution, $\text{CH}_3\text{COOH}$ , 2 M, 60 mL | Funnel                    |
| Sodium bicarbonate, $\text{NaHCO}_3$ , 10.5 g               | Graduated cylinder, 25-mL |
| Balloons, 6                                                 | Spatula                   |
| Erlenmeyer flasks, 125-mL, 6                                | Weighing dishes, 6        |

### Safety Precautions

*Acetic acid is a skin and eye irritant. Avoid contact with skin and eyes. Wear chemical splash goggles, chemical-resistant gloves, and a chemical-resistant apron. Please review current Material Safety Data Sheets for additional safety, handling, and disposal information.*

### Preparation

- Label six Erlenmeyer flasks 1–6. Using a graduated cylinder, add 10 mL of 2 M acetic acid to each flask.
- Obtain six weighing dishes and label them 1–6. Measure the appropriate amount of sodium bicarbonate into each weighing dish, as shown below.

| Sample                | 1      | 2      | 3      | 4      | 5      | 6      |
|-----------------------|--------|--------|--------|--------|--------|--------|
| Mass $\text{NaHCO}_3$ | 0.50 g | 1.00 g | 1.50 g | 2.00 g | 2.50 g | 3.00 g |

- Obtain six balloons. Stretch the balloons and blow them up at least once. Then let out as much air as possible.
- Use a powder funnel to add the first sodium bicarbonate sample (#1) to one of the balloons.
- Flatten out the balloon and then carefully place the neck of the balloon over the mouth of flask #1. Do not allow the solid to drop into the flask at this time.
- Repeat steps 4 and 5 using the remaining sodium bicarbonate samples in flasks #2–6.

### Procedure

- Line up the flasks on the demonstration table. Ask students to predict what will happen when the sodium bicarbonate is added to the acetic acid in each flask. How will the size of the inflated balloons vary when the reactions are complete?
- Lift each balloon in turn and shake it to allow the solid to fall into the solution. Make sure the neck of the balloon stays firmly attached to the flask.
- The reactions will be immediate and vigorous. The white solids will dissolve, the solutions will start to bubble and flow, and the balloons will become inflated.

# Stoichiometry Balloon Races Lab Answers

**AW Chickering**



**Stoichiometry Balloon Races Lab Answers:**

*Index Medicus* ,2001

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