

Stoichiometry Limiting Reagent

Answers

- 1 Consider the following reaction:



- (a) How many grams of NH_3 can be produced from the reaction of 28 g of N_2 and 25 g of H_2 ?

$$28 \text{ g N}_2 \times \frac{1 \text{ mol N}_2}{28 \text{ g N}_2} \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} \times \frac{17 \text{ g NH}_3}{1 \text{ mol NH}_3} = 34 \text{ g NH}_3$$

$$25 \text{ g H}_2 \times \frac{1 \text{ mol H}_2}{2 \text{ g H}_2} \times \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \times \frac{17 \text{ g NH}_3}{1 \text{ mol NH}_3} = 142 \text{ g NH}_3$$

34 g of NH_3 can be produced.

- (b) How much of the excess reagent is left over?

$$28 \text{ g N}_2 \times \frac{1 \text{ mol N}_2}{28 \text{ g N}_2} \times \frac{3 \text{ mol H}_2}{1 \text{ mol N}_2} \times \frac{2 \text{ g H}_2}{1 \text{ mol H}_2} = 6 \text{ g H}_2 \text{ used}$$

$$25 \text{ g H}_2 - 6 \text{ g H}_2 = 19 \text{ g H}_2 \text{ remains}$$

- 2 Consider the following reaction:



What volume of H_2 at STP is produced from the reaction of 50 g of Mg and 75 g of HCl ?

$$50 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.3 \text{ g Mg}} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = 46 \text{ L H}_2$$

$$75 \text{ g HCl} \times \frac{1 \text{ mol HCl}}{36.5 \text{ g HCl}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol HCl}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} = 23 \text{ L H}_2$$

23 L H_2 is produced.

- 3 Zinc and sulfur react to form zinc sulfide according to the following equation:



Suppose 25 g of zinc and 30 g of sulfur react. Which reagent is limiting?

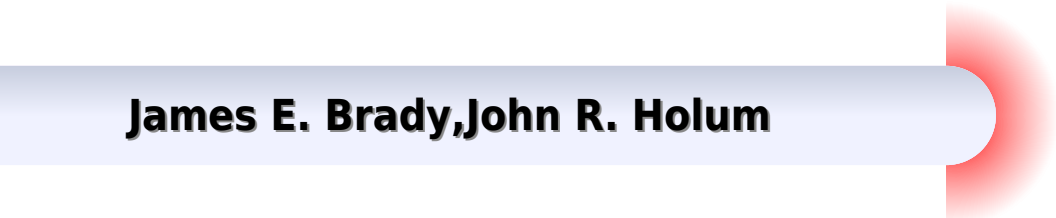
$$25 \text{ g Zn} \times \frac{1 \text{ mol Zn}}{65.4 \text{ g Zn}} \times \frac{1 \text{ mol ZnS}}{1 \text{ mol Zn}} \times \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 37.3 \text{ g ZnS}$$

$$30 \text{ g S} \times \frac{1 \text{ mol S}}{32 \text{ g S}} \times \frac{1 \text{ mol ZnS}}{1 \text{ mol S}} \times \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 91.2 \text{ g ZnS}$$

Zn is the limiting reagent.

Stoichiometry And Limiting Reactant Review

James E. Brady, John R. Holum



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Table of Contents Stoichiometry And Limiting Reactant Review

1. Understanding the eBook Stoichiometry And Limiting Reactant Review
 - The Rise of Digital Reading Stoichiometry And Limiting Reactant Review
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry And Limiting Reactant Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry And Limiting Reactant Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry And Limiting Reactant Review
 - Personalized Recommendations
 - Stoichiometry And Limiting Reactant Review User Reviews and Ratings
 - Stoichiometry And Limiting Reactant Review and Bestseller Lists
5. Accessing Stoichiometry And Limiting Reactant Review Free and Paid eBooks
 - Stoichiometry And Limiting Reactant Review Public Domain eBooks
 - Stoichiometry And Limiting Reactant Review eBook Subscription Services
 - Stoichiometry And Limiting Reactant Review Budget-Friendly Options
6. Navigating Stoichiometry And Limiting Reactant Review eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry And Limiting Reactant Review Compatibility with Devices
 - Stoichiometry And Limiting Reactant Review Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry And Limiting Reactant Review
 - Highlighting and Note-Taking Stoichiometry And Limiting Reactant Review
 - Interactive Elements Stoichiometry And Limiting Reactant Review
8. Staying Engaged with Stoichiometry And Limiting Reactant Review

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Stoichiometry And Limiting Reactant Review
- 9. Balancing eBooks and Physical Books Stoichiometry And Limiting Reactant Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry And Limiting Reactant Review
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry And Limiting Reactant Review
 - Setting Reading Goals Stoichiometry And Limiting Reactant Review
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry And Limiting Reactant Review
 - Fact-Checking eBook Content of Stoichiometry And Limiting Reactant Review
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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