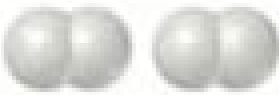


Stoichiometric Calculations

Equation:	$2 \text{H}_2(\text{g})$	+	$\text{O}_2(\text{g})$	$\longrightarrow$	$2 \text{H}_2\text{O}(\text{l})$
Molecules:	2 molecules H_2	+	1 molecule O_2	$\longrightarrow$	2 molecules H_2O
					
Mass (amu):	4.0 amu H_2	+	32.0 amu O_2	$\longrightarrow$	36.0 amu H_2O
Amount (mol):	2 mol H_2	+	1 mol O_2	$\longrightarrow$	2 mol H_2O
Mass (g):	4.0 g H_2	+	32.0 g O_2	$\longrightarrow$	36.0 g H_2O

The coefficients in the balanced equation give the ratio of *moles* of reactants and products

Stoichiometric Calculations Answers

Stanley E. Manahan



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