Solution Stoichiometry

1. What volume of a 0.35H AgNog W required to completely react with cs mi of a o dith Na U colution?

Ag NO3 + Na CI - Aga + Na Na

Required: V of Aq DO3

Hart with the other western

0-29 mol Nad 0.055 L Had last As Nos tolution: 1 % HARM LOWING I'M then charge

0.35 mal Ag DO3

= 0.03771 = 37.7 ml

2.37ml of a 0.29M &I column is required to completely reach with 42 ml of a Pb (NO3) e column. What was the original concentration of the lead (II) Vitrale bolution?

Ph (NO3)2 + 2 kI > PhI2 + 2 k NO3 Distant Required: 12 87 ml KI cln original concentration 12 0.2911 E2 The of the lead My D. DAH EI TIN M 12 ml Pb (NDab

0.29 mol k1 . 0.027 L . 1 mol Pb(10)2 . 1 2 mol kI D. Mall-= 0.1277

2 (42) M2 = 1(0.24)(37) M2 = 10.70 /89 H2 = 0.1277 /1-

3. What mass of solid Hagnessum Audmoxide can be produced if 15 mL of a D. 43 M Mg (NO2) 2 colution reacts completely with excess Nao4?

ING(NOTE+ 2 NOTH > Mg (DA)2 + 2 NOND3

pequired: man of cold Haynesium Hydroxide if time of a 0.63M Mg (1003)2

Given: 15 ml Ha (NO2)2, 0.43H Ha (NO2)2

NACL I Had 1-63 mol kg(NO3)2 0.0952 . [mol kg(0H) [mol Hq (1003)2

1 mol 49(0H)2 = 1659 49 (OH)2

1. What mass of solid Calcium Phosphate can be produced if 73ml of a 0.17M todium thoughte solution is mixed with 58ml of a D. 21 H calcium chloride solution?

2 Na , Poy + 3 Calls -> G Nacl + Cas (Pay)2

biven:

Town Nay Pod Seret Calla D.M. Was Pos O. HH Calla

Required: gio. leg/ mass (g) cas (PM)2 From: Has Pos

D.H mol Nag Pay . 0.0721 . Inol Crollage

. 310. 189 (43/00) = 1.925g (43 PD) mol (azlifor)

Fram Caula 0.0581 Irul cas (800)2. 3/0 9 cas (800) Ima 2 mol cays

= [1.4381 a caz (004)e],

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ML Yell

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