

Stoichiometry Review

Directions: Each of the problems below reviews a different type of stoichiometry. Be sure to read the problem carefully to know what it is asking for.

Mass-to-Mass Problems

1. How many grams of potassium chloride, KCl, are produced during the decomposition of 100.0 grams of potassium chlorate, KClO_3 ?



$$\frac{100.0 \text{ g KClO}_3}{1} \times \frac{1 \text{ mol KClO}_3}{122.545 \text{ g KClO}_3} \times \frac{1 \text{ mol KCl}}{1 \text{ mol KClO}_3} \times \frac{74.548 \text{ g KCl}}{1 \text{ mol KCl}} = \underline{60.83 \text{ g KCl}}$$

2. Find the number of grams of HCl needed to react completely with 0.50 moles of magnesium.



$$\frac{0.5 \text{ mol Mg}}{1} \times \frac{1 \text{ mol HCl}}{1 \text{ mol Mg}} \times \frac{36.458 \text{ g HCl}}{1 \text{ mol HCl}} = \underline{18.23 \text{ g HCl}}$$

3. Find the mass in grams of hydrogen gas produced when 14.0 moles of HCl are added to an excess amount of magnesium.



$$\frac{14 \text{ mol HCl}}{1} \times \frac{1 \text{ mol H}_2}{2 \text{ mol HCl}} \times \frac{2.016 \text{ g H}_2}{1 \text{ mol H}_2} = \underline{14.112 \text{ g H}_2}$$

4. If 0.75 moles of ammonia are needed, how many grams of nitrogen will be consumed?



$$\frac{0.75 \text{ mol NH}_3}{1} \times \frac{1 \text{ mol N}_2}{2 \text{ mol NH}_3} \times \frac{28.014 \text{ g N}_2}{1 \text{ mol N}_2} = \underline{10.51 \text{ g N}_2}$$

Stoichiometry Review Key

John Moore, Richard H. Langley

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the authors' names.

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