

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) =14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) +4(2) =14 O atoms

b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) =6 O atoms

Product side

4 Fe atoms

2(3) =6 O atoms

b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃:4 mol Fe

Stoichiometry Problems In Chemistry And Answers

Paul A Krieger



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heart of chemistry lies the concept of atoms. Atoms are the fundamental units of matter and they consist of a nucleus containing protons and neutrons surrounded by electrons. Each element on the periodic table is defined by the number of protons in its nucleus. So elements are like a unique fingerprint for the matter. Exactly. There are over a hundred known elements and they combine in various ways to form compounds. These compounds can be as simple as water H_2O or as complex as DNA. I've heard of the periodic table. Is it like a cheat sheet for chemistry? You got it. The periodic table organizes elements based on their properties and atomic numbers. It helps us understand the relationships between different elements and their behaviors in chemical reactions. Chemical reactions. That sounds exciting. How do they work? Chemical reactions occur when atoms rearrange to form new substances with different properties. It's like a dance of atoms where they break old bonds and create new ones. I can visualize that. Are there different types of chemical reactions? Indeed. There are various types such as synthesis, decomposition, single replacement, and double replacement reactions. Each type has unique characteristics and applications. Chemistry seems like a puzzle with all these pieces coming together. That's a great analogy. And speaking of puzzles, we have to mention acids and bases. They play a vital role in chemistry and can be found in many everyday substances. I've heard of acids and bases. They have something to do with the pH scale, right? Exactly. The pH scale measures the acidity or basicity of a substance. It ranges from 0 to 14, with 7 being neutral. Substances with a pH less than 7 are acidic, while those with a pH greater than 7 are basic. I'm starting to see how chemistry is all around us, shaping the world we live in. Absolutely. Chemistry influences everything from the food we eat to the medicines we take, the materials we use, and even the air we breathe. It's a central science that connects many other fields of study. This is truly captivating. I can't wait to learn more about the wonders of chemistry and delve into its intricacies. I'm thrilled that you're so enthusiastic. In the upcoming chapters, we'll explore topics like thermodynamics, organic chemistry, chemical kinetics, and even the cutting-edge world of quantum chemistry. I'm ready for the adventure. Let's uncover the mysteries of the fascinating world of chemistry together. You bet. Get ready for an amazing journey into the heart of matter and the magic of molecules. Chemistry awaits.

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