

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) = 14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) + 4(2) = 14 O atoms

- b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) = 6 O atoms

Product side

4 Fe atoms

2(3) = 6 O atoms

- b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃: 4 mol Fe

Stoichiometry Problems Answers

Dilip K. Das, Rajaram K. Prabhudesai

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Table of Contents Stoichiometry Problems Answers

1. Understanding the eBook Stoichiometry Problems Answers
 - The Rise of Digital Reading Stoichiometry Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Problems Answers
 - Personalized Recommendations
 - Stoichiometry Problems Answers User Reviews and Ratings
 - Stoichiometry Problems Answers and Bestseller Lists

5. Accessing Stoichiometry Problems Answers Free and Paid eBooks
 - Stoichiometry Problems Answers Public Domain eBooks
 - Stoichiometry Problems Answers eBook Subscription Services
 - Stoichiometry Problems Answers Budget-Friendly Options
6. Navigating Stoichiometry Problems Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Problems Answers Compatibility with Devices
 - Stoichiometry Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Problems Answers
 - Highlighting and Note-Taking Stoichiometry Problems Answers
 - Interactive Elements Stoichiometry Problems Answers
8. Staying Engaged with Stoichiometry Problems Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Problems Answers
9. Balancing eBooks and Physical Books Stoichiometry Problems Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Problems Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Problems Answers
 - Setting Reading Goals Stoichiometry Problems Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Problems Answers
 - Fact-Checking eBook Content of Stoichiometry Problems Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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