

Stoichiometry with Percent Yield

4) In manufacturing steel, iron oxide reacts with carbon to form iron and carbon monoxide as described by the reaction below. What is the percent yield of iron if 1400.3 grams of iron is produced from the reaction of 4.128 kg grams of Fe_2O_3 and excess carbon? The reaction is $\text{Fe}_2\text{O}_3 + \text{C} \rightarrow 2\text{Fe} + \text{CO}$

5) 0.00003 grams of Barium-Chloride react with excess Sodium Sulfate. What is the percentage yield of Barium Sulfate if 0.000280 grams of BaSO_4 is produced? $\text{BaCl}_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$

6) 0.45005 grams of iron (II) sulfide reacts with nitrogen in the reaction below. If 0.58252 grams formed in the reaction, what is the percent yield? $2\text{FeS} + \text{N}_2 \rightarrow 4\text{Fe} + 2\text{N}_2\text{S}$

Name _____ Section _____ Date _____

Stoichiometry and Percent yield

Directions: Determine the amount of product that can be produced and the percent yield of each reaction.

1) 3.81 grams of H_3PO_4 react with excess NaOH in the following reaction: $\text{H}_3\text{PO}_4 + 3\text{NaOH} \rightarrow \text{Na}_3\text{PO}_4 + 3\text{H}_2\text{O}$. What is the percent yield if 3.05 grams of water are produced?

2) 9.239 grams of iron are produced when aluminum reacts with excess iron oxide in the reaction below. What is the percent yield if 2.95 grams of aluminum react? $\text{Fe}_2\text{O}_3 + \text{Al} \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$

3) What is the percentage yield of CO_2 formed when 12.47 grams of CO_2 is collected after the reaction of 34.80 grams of C_2H_6 in the reaction: $\text{C}_2\text{H}_6 + \text{SO}_2 \rightarrow \text{SOO}_2 + \text{CO}_2$

7) What is the percent yield when 5.552 grams of PbO_2 is produced in the reaction of 12.52 grams of Pb and excess $\text{Pb}(\text{NO}_3)_2$ in the reaction $\text{Pb}(\text{NO}_3)_2 + 2\text{Pb} \rightarrow \text{PbO}_2 + 2\text{PbNO}_2$

8) 21.28 grams of H_2S combine with excess oxygen. What is the percent yield if 14.48 grams of SO_2 are formed? $2\text{H}_2\text{S} + \text{SO}_2 \rightarrow 2\text{SO}_2 + 2\text{H}_2\text{O}$

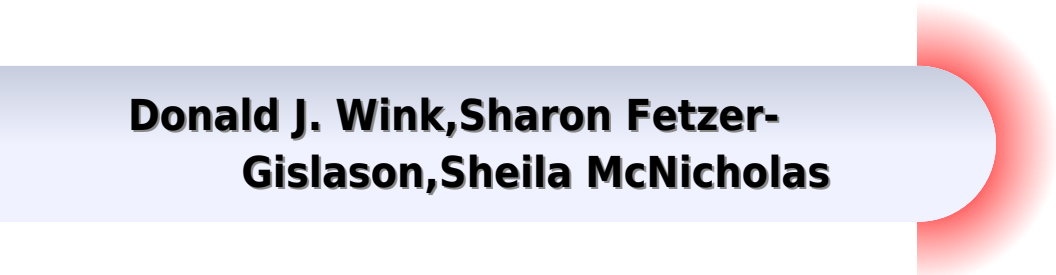
9) When 13.3 grams of $\text{Ba}(\text{OH})_2$ reacts with excess $\text{HC}_2\text{H}_3\text{O}_2$ 6.08 grams of $\text{Ba}(\text{C}_2\text{H}_3\text{O}_2)_2$ are formed. What is the percent yield for this reaction? $\text{Ba}(\text{OH})_2 + 2\text{HC}_2\text{H}_3\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{Ba}(\text{C}_2\text{H}_3\text{O}_2)_2$

10) 95.805 grams of sulfur combine with excess oxygen and water to form 77.773 grams of sulfuric acid. What is the percent yield? In the reaction: $\text{S} + \text{SO}_2 + 2\text{H}_2\text{O} \rightarrow 2\text{H}_2\text{SO}_4$

A Chemistry Worksheet
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Stoichiometry Percent Yield Answers

**Donald J. Wink, Sharon Fetzer-
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Philosophy of "Packer" Pedagogy: Vince Lombardi, Critical Thinking and Problem-Based Learning, 2nd Edition David Pushkin, 2011-09-07 Vince Lombardi the greatest head coach in professional football history began his journey towards greatness as a high school chemistry and physics teacher in 1939 The core principles he developed for eight years in an Englewood New Jersey science classroom helped provide the foundation for his legendary Green Bay Packers winning five NFL championship titles during a seven year period 1961-1967 Dave Pushkin a former football player turned chemistry and physics professor for 25 years was greatly influenced by Lombardi's core principles developing his own teaching practices from these principles as well as cognitive and curricular theories This book presents an intellectual merging between Lombardi's principles with respect to football and educational theory philosophy

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