

Mole to Mole Stoichiometry Worksheet

- 1) In the reaction below, how many moles of aluminum will be used when reacted with 1.33 moles of oxygen?
 $4 \text{ Al} + 3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$

$$1.33 \text{ mol O}_2 \times \frac{4 \text{ mol Al}}{3 \text{ mol O}_2} = \boxed{1.77 \text{ mol Al}}$$

- 2) How many moles of hydrogen will be produced when reacted with 0.0204 moles of sodium in the reaction?
 $2 \text{ Na} + 2 \text{ H}_2\text{O} \rightarrow 2 \text{ NaOH} + \text{H}_2$

$$0.0204 \text{ mol Na} \times \frac{1 \text{ mol H}_2}{2 \text{ mol Na}} = \boxed{0.0102 \text{ mol H}_2}$$

- 3) Using the equation $2 \text{ Al(NO}_3)_3 + 3 \text{ MgBr}_2 \rightarrow 3 \text{ Mg(NO}_3)_2 + 2 \text{ AlBr}_3$
 a. Find the number of moles $\text{Al(NO}_3)_3$ used to produce 2.3 moles of $\text{Mg(NO}_3)_2$

$$2.3 \text{ mol Mg(NO}_3)_2 \times \frac{2 \text{ mol Al(NO}_3)_3}{3 \text{ mol Mg(NO}_3)_2} = \boxed{1.53 \text{ mol Al(NO}_3)_3}$$

- b. Find the number of moles $\text{Al(NO}_3)_3$ used to produce 0.78 moles of AlBr_3

$$0.78 \text{ mol AlBr}_3 \times \frac{2 \text{ mol Al(NO}_3)_3}{2 \text{ mol AlBr}_3} = \boxed{0.78 \text{ mol Al(NO}_3)_3}$$

- c. Find the number of moles MgBr_2 that reacts with 1.05 moles of $\text{Al(NO}_3)_3$

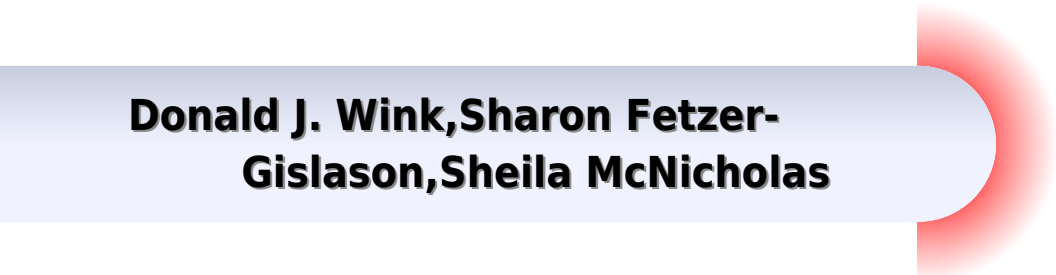
$$1.05 \text{ mol Al(NO}_3)_3 \times \frac{3 \text{ mol MgBr}_2}{2 \text{ mol Al(NO}_3)_3} = \boxed{1.58 \text{ mol MgBr}_2}$$

- 4) In the reaction below, how many moles of carbon dioxide are produced if 19.65 moles of water are produced?
 $2 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$

$$19.65 \text{ moles H}_2\text{O} \times \frac{4 \text{ mol CO}_2}{6 \text{ mol H}_2\text{O}} = \boxed{13.1 \text{ mol CO}_2}$$

Stoichiometry Mole Practice Problems Answer Key

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ChemDiscovery Teacher Edition Olga I. Agapova, 2002 Basics for Chemistry David A. Ucko, 2013-09-24 Basics of Chemistry provides the tools needed in the study of General Chemistry such as problem solving skills calculation methods and the language and basic concepts of chemistry The book is designed to meet the specific needs of underprepared students Concepts are presented only as they are needed and developed from the simple to the complex The text is divided into 18 chapters each covering some particular aspect of chemistry such as matter energy and measurement the properties of atoms description of chemical bonding study of chemical change and nuclear and organic chemistry Undergraduate students will find the book as a very valuable academic material *The Practice of Chemistry* Donald J. Wink, Sharon

Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB **Introduction to**

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electrochemical engineering environmental engineering safety engineering and industrial chemistry The chief objective of this text is to prepare students to make analysis of chemical processes through calculations and also to develop in them systematic problem solving skills The students are introduced not only to the application of law of combining proportions to chemical reactions as the word stoichiometry implies but also to formulating and solving material and energy balances in processes with and without chemical reactions The book presents the fundamentals of chemical engineering operations and processes in an accessible style to help the students gain a thorough understanding of chemical process calculations It also covers in detail the background materials such as units and conversions dimensional analysis and dimensionless groups property estimation P V T behaviour of fluids vapour pressure and phase equilibrium relationships humidity and saturation With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations Key Features SI units are used throughout the book Presents a thorough introduction to basic chemical engineering principles Provides many worked out examples and exercise problems with answers Objective type questions included at the end of the book serve as useful review material and also assist the students in preparing for competitive examinations such as GATE *MCAT General Chemistry Review 2018-2019* Kaplan Test Prep, 2017-07-04 Kaplan's MCAT Complete 7 Book Set Subject Review has all the information and strategies you need to score higher on the MCAT These books feature more practice than any other guide plus targeted strategy review opportunities for self analysis and thorough information on all of the critical thinking skills necessary for MCAT success from the creators of the 1 MCAT prep course From publisher's description

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