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Date: _____

Student Exploration: Stoichiometry

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, conversion factor, dimensional analysis, molar mass, mole, molecular mass, stoichiometry

Activity A: Moles

Get the Gizmo ready:

- Check that the equation is still:
$$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$$
- If not, click **New equation** until it reappears.



Question: How do scientists find the molecular mass and molar mass of a substance?

- Calculate:** The **molecular mass** of a molecule is the sum of the masses of each atom in the molecule. The unit of molecular mass is the universal mass unit (u).

Iron's atomic mass is 55.85 u, carbon's mass is 12.01 u, and oxygen's mass is 16.00 u.

- Calculate the molecular mass of carbon monoxide (CO) by adding the atomic mass of carbon and the atomic mass of oxygen: 28.01g
- Calculate the molecular mass of carbon dioxide (CO₂): 44.01g
- Calculate the molecular mass of iron (III) oxide (Fe₂O₃): 159.7 g

- Infer:** A mole of a substance has a mass in grams that is equal to the molecular mass. For example, a carbon atom has a mass of 12.01 u. A mole of carbon has a mass of 12.01 g. Based on their molecular masses, list the **molar mass** of each substance:

Fe₂O₃ 159.7 g CO 28.01g Fe 55.85g CO₂ 44.01g

Check your answers on the Gizmo by inspecting the middle row of tiles on the right side of the Gizmo. These tiles show the units "1 mol" on top and "g" below.

- Practice:** Hydrogen has an atomic mass of 1.01 u. What is the molar mass of these substances?

H₂O 18.02g CH₄ 16.05g H₂CO₃ 62.03g C₆H₁₂O₆ 180.18g

Stoichiometry Explore Learning Activity B

R Bogdan



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Table of Contents Stoichiometry Explore Learning Activity B

1. Understanding the eBook Stoichiometry Explore Learning Activity B
 - The Rise of Digital Reading Stoichiometry Explore Learning Activity B
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Explore Learning Activity B
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Explore Learning Activity B
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Explore Learning Activity B
 - Personalized Recommendations
 - Stoichiometry Explore Learning Activity B User Reviews and Ratings
 - Stoichiometry Explore Learning Activity B and Bestseller Lists

5. Accessing Stoichiometry Explore Learning Activity B Free and Paid eBooks
 - Stoichiometry Explore Learning Activity B Public Domain eBooks
 - Stoichiometry Explore Learning Activity B eBook Subscription Services
 - Stoichiometry Explore Learning Activity B Budget-Friendly Options
6. Navigating Stoichiometry Explore Learning Activity B eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Explore Learning Activity B Compatibility with Devices
 - Stoichiometry Explore Learning Activity B Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Explore Learning Activity B
 - Highlighting and Note-Taking Stoichiometry Explore Learning Activity B
 - Interactive Elements Stoichiometry Explore Learning Activity B
8. Staying Engaged with Stoichiometry Explore Learning Activity B
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Explore Learning Activity B
9. Balancing eBooks and Physical Books Stoichiometry Explore Learning Activity B
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Explore Learning Activity B
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Explore Learning Activity B
 - Setting Reading Goals Stoichiometry Explore Learning Activity B
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Explore Learning Activity B
 - Fact-Checking eBook Content of Stoichiometry Explore Learning Activity B
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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